

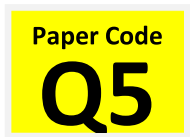
# MENIIT

NEET | IIT-JEE | FOUNDATION

Corporate Office: 44-A/1, Kalu Sarai, New Delhi 110016 | Web: [www.meniit.com](http://www.meniit.com)

Maximum Marks: 720

Time: 3 Hours 20 Minutes



## NEET (UG) – 2022

### IMPORTANT INSTRUCTIONS

1. The Answer Sheet is inside this Test Booklet. When you are directed to open the Test Booklet, take out the Answer Sheet and fill in the particulars on **side-1** and **side-2** carefully with blue/black ball point pen only.
2. The test is of 3 hours 20 minutes duration and Test Booklet contains **200** multiple-choice questions (four option with a single correct answer) from Physics, Chemistry and Biology (Botany and Zoology). 50 questions in each subject are divided into two sections (A and B) as per details given below:
  - (a) **Section A** shall consist of 35 (Thirty-five) Questions in each subject (Question Nos 1 to 35, 51 to 85, 101 to 135 and 151 to 185). All questions are compulsory.
  - (b) **Section B** shall consist of 15 (Fifteen) questions in each subject (Question Nos – 36 to 50, 86 to 100, 136 to 150 and 186 to 200). In section B, a candidate needs to attempt any 10 (Ten) questions out of 15 (Fifteen) in each subject.

Candidates are advised to read all 15 questions in each subject of section B before they start attempting the question paper. In the event of a candidate attempting more than ten questions, the first ten questions answered by the candidate shall be evaluated.
3. Each question carries 4 marks. For each correct response, the candidate will get 4 marks. For each incorrect response, one mark will be deducted from the total scores. The maximum marks are 720.
4. Use Blue/Black Ball Point Pen Only for writing particulars on this page/markings responses on Answer Sheet.
5. Rough work is to be done on the space provided for this purpose in the Test Booklet only.
6. **On completion of the test, the candidate must hand over the Answer Sheet (ORIGINAL and OFFICE Copy) to the invigilator before leaving the Room/Hall. The candidates are allowed to take away this Test Booklet with them.**
7. The CODE for this Booklet is **Q5**. Make sure that the CODE printed on Original Copy of the Answer Sheet is the same as on this Test Booklet. In case of discrepancy, the candidate should immediately report the matter to the Invigilator for replacement of both the Test Booklet and the Answer Sheet.
8. The candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your Roll No. anywhere else except in the specified space in the Test Booklet/Answer Sheet.
9. Use of white fluid for correction is **NOT** permissible on the Answer Sheet.
10. Each candidate must show on demand his/her Admit Card to the Invigilator.
11. No candidate, without special permission of the Superintendent or Invigilator, would leave his/her seat.
12. The candidates should not leave the Examination Hall without handing over their Answer Sheet to the Invigilator on duty and sign the Attendance Sheet twice. **Cases where a candidate has not signed the Attendance Sheet second time will be deemed not to have handed over the Answer Sheet and dealt with as an unfair means case.**
13. Use of Electronic/Manual Calculator is prohibited.
14. The candidates are governed by all Rules and Regulations of the examination with regard to their conduct in the Examination Hall. All cases of unfair means will be dealt with as per Rules and Regulations of this Examination.
15. No part of the Test Booklet and Answer Sheet shall be detached under any circumstances.
16. The candidates will write the Correct Test Booklet Code as given in the Test Booklet/Answer Sheet in the Attendance Sheet.
17. Compensatory time of one hour five minutes will be provided for the examination of three hours and 20 minutes duration, whether such candidate (having a physical limitation to write) uses the facility of scribe or not.

## SECTION - A (PHYSICS)

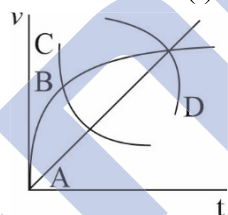
1. A square loop of side 1 m and resistance  $1 \Omega$  is placed in a magnetic field of 0.5 T. If the plane of loop is perpendicular to the direction of a magnetic field, the magnetic flux through the loop is
- (1) 2 weber                      (2) 0.5 weber                      (3) 1 weber                      (4) zero weber
2. When light propagates through a material medium of relative permittivity  $\epsilon_r$  and relative permeability  $\mu_r$ , the velocity of light,  $v$  is given by : ( $c$  – velocity of light in vacuum)

(1)  $v = c$                       (2)  $v = \sqrt{\frac{\mu_r}{\epsilon_r}}$                       (3)  $v = \sqrt{\frac{\epsilon_r}{\mu_r}}$                       (4)  $v = \frac{c}{\sqrt{\epsilon_r \mu_r}}$

3. When two monochromatic lights of frequency,  $\nu$  and  $\frac{\nu}{2}$  are incident on a photoelectric metal, their stopping potential becomes  $\frac{V_s}{2}$  and  $V_s$  respectively. The threshold frequency for this metal is

(1)  $2\nu$                       (2)  $3\nu$                       (3)  $\frac{2}{3}\nu$                       (4)  $\frac{3}{2}\nu$

4. A spherical ball is dropped in a long column of a highly viscous liquid. The curve in the graph shown, which represents the speed of the ball ( $v$ ) as a function of time ( $t$ ) is:



- (1) A                      (2) B                      (3) D                      (4) D

5. Given below are two statements:

Statements I:

Biot–Savart’s law gives us the expression for the magnetic field strength of an infinitesimal current element ( $Idl$ ) of a current carrying conductor only

Statements II:

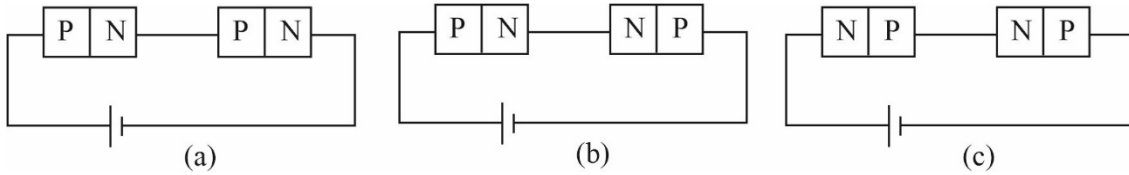
Biot–Savart’s law is analogous to Coulomb’s inverse square law of charge  $q$ , with the former being related to the field produced by a scalar source,  $Idl$  while the latter being produced by a vector source,  $q$ .

In light of above statements choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct  
 (2) Both statement I and Statement II are incorrect  
 (3) Statement I is correct and Statement II is incorrect  
 (4) Statement I is incorrect and Statement II is correct
6. As the temperature increases, the electric resistance:
- (1) increases for both conductors and semiconductors  
 (2) decreases for both conductors and semiconductors  
 (3) increases for conductors but decreases for semiconductors  
 (4) decrease for conductors but increases for semiconductor.

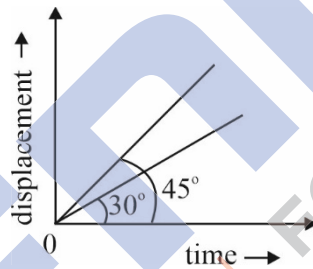
7. Two resistors of resistance,  $100\ \Omega$  and  $200\ \Omega$  are connected in parallel in an electrical circuit. The ratio for the thermal energy developed in  $100\ \Omega$  to that in  $200\ \Omega$  in a given time is :
- (1) 1 : 2                      (2) 2 : 1                      (3) 1 : 4                      (4) 4 : 1

8.



In the given circuit (a), (b) and (c), the potential drop across the two p–n junctions are equal in :

- (1) Circuit (a) only    (2) Circuit (b) only    (3) Circuit (c) only    (4) Circuit (a) only (c)
9. The peak voltage of the ac source is equal to :
- (1) the value of voltage supplied to the circuit    (2) the rms value of the ac source
- (3)  $\sqrt{2}$  times the rms value of the ac source    (4)  $\frac{1}{\sqrt{2}}$  times the rms value of the ac source
10. The displacement–times graphs of two moving particles make angle of  $30^\circ$  and  $45^\circ$  with the x–axis as shown in the figure. The ratio of their respective velocity is :



- (1)  $\sqrt{3}:1$                       (2) 1 : 1                      (3) 1 : 2                      (4)  $1:\sqrt{3}$
11. The angel between the electric lines of forces and the equipotential surface is :
- (1)  $0^\circ$                       (2)  $45^\circ$                       (3)  $90^\circ$                       (4)  $180^\circ$
12. The dimensions  $[ML^{-2}A^{-2}]$  belong to the:
- (1) magnetic flux                      (2) self inductance
- (3) magnetic permeability                      (4) electric permittivity
13. If a soap bubble expands, the pressure inside the bubble :
- (1) decreases                      (2) increases
- (3) remains the same                      (4) is equal to the atmospheric pressure
14. The energy that will be ideally radiated by a 100 kW transmitter in 1 hour is :
- (1)  $36 \times 10^7 J$                       (2)  $36 \times 10^4 J$                       (3)  $36 \times 10^5 J$                       (4)  $1 \times 10^5 J$
15. In half wave rectification, if the input frequency is 60 Hz, then the output frequency would be :
- (1) zero                      (2) 30 Hz                      (3) 60 Hz                      (4) 120 Hz
16. Two objects of mass 10 kg and 20 kg respectively are connected to the two ends of a rigid rod of length 10 m with negligible mass. The distance of the centre of mass of the system from the 10 kg mass is :
- (1)  $\frac{10}{3}m$                       (2)  $\frac{20}{3}m$                       (3) 10m                      (4) 5m

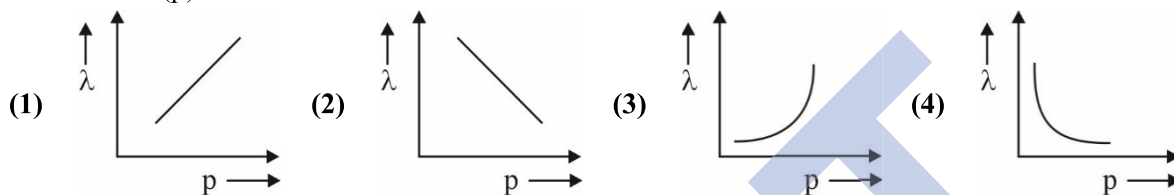
17. Match List-I with List-II :

	<b>List-I (Electromagnetic waves)</b>		<b>List-II (Wavelength)</b>
(a)	AM radio waves	(i)	$10^{-10}$ m
(b)	Microwaves	(ii)	$10^2$ m
(c)	Infrared radiations	(iii)	$10^{-2}$ m
(d)	X-rays	(iv)	$10^{-4}$ m

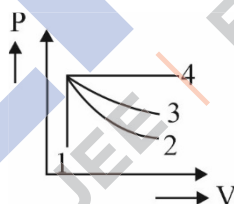
Choose the correct answer from the options given below:

- (1) (a)–(iv), (b)–(iii), (c) – (ii), (d) – (i)      (2) (a)–(iii), (b)–(ii), (c) – (i), (d) – (iv)  
 (3) (a)–(iii), (b)–(iv), (c) – (ii), (d) – (i)      (4) (a)–(ii), (b)–(iii), (c) – (iv), (d) – (i)
18. An electric lift with a maximum load of 2000 kg (lift + passenger) is moving up with a constant speed  $1.5 \text{ ms}^{-1}$ . The frictional force opposing the motion is 2000 N. The minimum power delivered by the motor to lift in watts is : ( $g = 10 \text{ ms}^{-2}$ )  
 (1) 23000      (2) 20000      (3) 34500      (4) 23500
19. In a Young's double slit experiment, a student observes 8 fringes in a certain segment of screen when a monochromatic light of 600 nm wavelength is used. If the wavelength of light is changed to 400 nm, then the number of fringes he would observe in the same region of the screen is:  
 (1) 6      (2) 8      (3) 9      (4) 12
20. Two hollow conducting spheres of radii  $R_1$  and  $R_2$  ( $R_1 \gg R_2$ ) have equal charge. The potential would be:  
 (1) more on bigger sphere      (2) more on smaller sphere  
 (3) equal on both the spheres      (4) dependent on the material property of the sphere.
21. In the given nuclear reaction, the elements X is:  
 ${}_{11}^{22}\text{Na} \rightarrow X + e^+ + \nu$   
 (1)  ${}_{11}^{23}\text{Na}$       (2)  ${}_{10}^{23}\text{Ne}$       (3)  ${}_{10}^{22}\text{Ne}$       (4)  ${}_{12}^{22}\text{Mg}$
22. The ratio of the radius of gyration of a thin uniform disc about an axis passing through its centre and normal to its plane to the radius of gyration of the disc about its diameter is :  
 (1) 2 : 1      (2)  $\sqrt{2} : 1$       (3) 4 : 1      (4)  $1 : \sqrt{2}$
23. Let  $T_1$  and  $T_2$  be the energy of an electron in the first and second excited states of hydrogen atom, respectively. According to Bohr's model of an atom, the ratio  $T_1 : T_2$  is :  
 (1) 1 : 4      (2) 4 : 1      (3) 4 : 9      (4) 9 : 4
24. A light ray falls on a glass surface of refractive index  $\sqrt{3}$ , at an angle  $60^\circ$ . The angle between the refracted and reflected rays would be:  
 (1)  $30^\circ$       (2)  $60^\circ$       (3)  $90^\circ$       (4)  $120^\circ$
25. A copper wire of length 10 m and radius  $(10^{-2} / \pi)$  m has electrical resistance of  $10\Omega$ . The current density in the wire for an electric field strength of  $10(\text{V/m})$  is:  
 (1)  $10^4 \text{ A/m}^2$       (2)  $10^6 \text{ A/m}^2$       (3)  $10^{-5} \text{ A/m}^2$       (4)  $10^5 \text{ A/m}^2$
26. A biconvex lens has radii of curvature, 20 cm each, If the refractive index of the material of the lens is 1.5, the power of the lens is:  
 (1) +2D      (2) +20D      (3) +5D      (4) infinity

27. A long solenoid of radius 1 mm has 100 turns per mm. if 1 A current flows in the solenoid, the magnetic field strength at the centre of the solenoid is:  
 (1)  $6.28 \times 10^{-2} \text{ T}$       (2)  $12.56 \times 10^{-2} \text{ T}$       (3)  $12.56 \times 10^{-4} \text{ T}$       (4)  $6.28 \times 10^{-4} \text{ T}$
28. A body of mass 600 g experiences a gravitational force of 3.0 N, when placed at a particular point. The magnitude of the gravitational field intensity at that point is :  
 (1) 0.05 N / kg      (2) 50 N / kg      (3) 20 N / kg      (4) 180 N / kg
29. The graph which shows the variation of the de-Broglie wavelength ( $\lambda$ ) of a particle and its associated momentum (p) is :



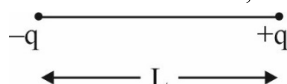
30. The ratio of the distances travelled by a freely falling body in the 1<sup>st</sup>, 2<sup>nd</sup>, 3<sup>rd</sup> and 4<sup>th</sup> second:  
 (1) 1 : 2 : 3 : 4      (2) 1 : 4 : 9 : 16      (3) 1 : 3 : 5 : 7      (4) 1 : 1 : 1 : 1
31. The angular speed of a fly wheel moving with uniform angular acceleration changes from 1200 rpm to 3120 rpm in 16 seconds. The angular acceleration in  $\text{rad/s}^2$  is:  
 (1)  $2\pi$       (2)  $4\pi$       (3)  $12\pi$       (4)  $104\pi$
32. An ideal gas undergoes four different processes from the same initial state as shown in the figure below: Those process are adiabatic, isothermal, isobaric and isochoric. The curve which represents the adiabatic process among 1, 2, 3 and 4 is:



- (1) 1      (2) 2      (3) 3      (4) 5
33. If the initial tension on a stretched string is doubled, then the ratio of the initial and final speeds of a transverse wave along the string is:  
 (1) 1 : 1      (2)  $\sqrt{2} : 1$       (3)  $1 : \sqrt{2}$       (4) 1 : 2
34. Plane angle and solid angle have:  
 (1) Units but no dimension      (2) Dimensions but not units  
 (3) No units and no dimensions      (4) Both units and dimensions
35. A shell of mass m is at rest initially. It explodes into three fragments having mass in the ratio 2 : 2 : 1. If the fragments having equal mass fly off along mutually perpendicular directions with speed v, the speed of the third (lighter) fragment is :  
 (1) v      (2)  $\sqrt{2}v$       (3)  $2\sqrt{2}v$       (4)  $3\sqrt{2}v$

## SECTION - B (PHYSICS)

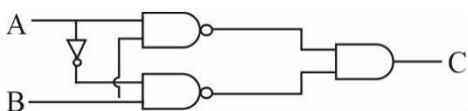
36. The area of rectangular field (in  $\text{m}^2$ ) of length 55.3 and breadth 25m after rounding off the value of correct significant digits is:  
 (1)  $138 \times 10^1$       (2) 1382      (3) 1382.5      (4)  $14 \times 10^2$
37. A big circular coil of 1000 turns, and average radius 10 m is rotating about its horizontal diameter at  $2 \text{ rad s}^{-1}$ . If the vertical component of earth's magnetic field at that place is  $2 \times 10^{-5} \text{ T}$  and electrical resistance of the coil is  $12.56 \Omega$ , then the maximum induced current in the coil will be:  
 (1) 0.25 A      (2) 1.5 A      (3) 1 A      (4) 2 A
38. Two point charges  $-q$  and  $+q$  are placed at a distance of  $L$ , as shown in the figure.



The magnitude of electric field intensity at a distance  $R$  ( $R \gg L$ ) varies as :

- (1)  $\frac{1}{R^2}$       (2)  $\frac{1}{R^3}$       (3)  $\frac{1}{R^4}$       (4)  $\frac{1}{R^6}$

39.



The truth table for the given logic circuit is:

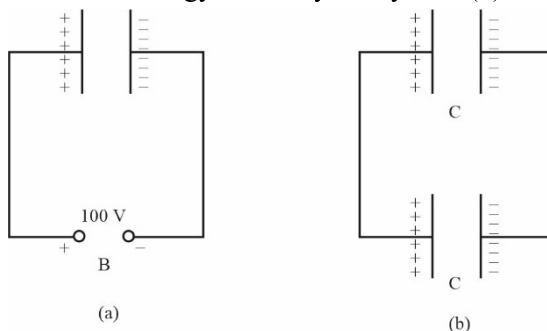
(1)	A	B	C
	0	0	0
	0	1	1
	1	0	1
	1	1	0

(2)	A	B	C
	0	0	1
	0	1	0
	1	0	0
	1	1	1

(3)	A	B	C
	0	0	1
	0	1	0
	1	0	1
	1	1	0

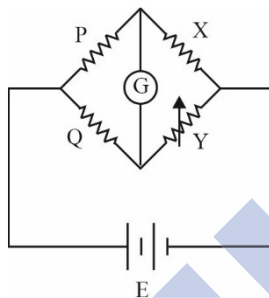
(4)	A	B	C
	0	0	0
	0	1	1
	1	0	0
	1	1	1

40. A capacitor of capacitance  $C = 900 \text{ pF}$  is charged fully by 100 V battery B as shown in figure (a). The it is disconnected from the battery and connected to another uncharged capacitor of capacitance  $C = 900 \text{ pF}$  as shown in figure (b). The electrostatic energy stored by the system (b) is :



- (1)  $4.5 \times 10^{-6} \text{ J}$       (2)  $3.25 \times 10^{-6} \text{ J}$       (3)  $2.25 \times 10^{-6} \text{ J}$       (4)  $1.5 \times 10^{-6} \text{ J}$

41. Two transparent media A and B are separated by a plane boundary. The speed of light in those media are  $1.5 \times 10^8$  m/s and  $2.0 \times 10^8$  m/s, respectively. The critical angle for a ray of light for these two media is  
 (1)  $\sin^{-1}(0.500)$       (2)  $\sin^{-1}(0.750)$       (3)  $\tan^{-1}(0.500)$       (4)  $\tan^{-1}(0.750)$
42. A ball is projected with a velocity,  $10 \text{ ms}^{-1}$ , at an angle of  $60^\circ$  with the vertical direction, its speed at the highest point of its trajectory will be :  
 (1) zero      (2)  $5\sqrt{3}\text{ms}^{-1}$       (3)  $5\text{ms}^{-1}$       (4)  $10\text{ms}^{-1}$
43. A Wheatstone bridge is used to determine the value of unknown resistance X by adjusting the variable resistance Y as shown in the figure. For the moving precise measurement of X, the resistance P and Q



- (1) should be approximately equal to  $2X$       (2) should be approximately equal and are small  
 (3) should be very large and unequal      (4) do not play any significant role
44. Given below are two statements : One is labelled as **Assertion (A)** and the other is labelled as **Reason (R)**.  
**Assertion (A) :**  
 The stretching of a spring is determined by the shear Modulus of the material of the spring.  
**Reason (R) :**  
 A coil spring of copper has more tensile strength than a steel spring of same dimensions.  
 In the light of the above statements, choose the most appropriate answer from the options given below :  
 (1) Both (A) and (R) are true and (R) is the correct explanation of (A)  
 (2) Both (A) and (R) are true and (R) is not the correct explanation of (A)  
 (3) (A) is true but (R) is false  
 (4) (A) is false but (R) is true
45. A series LCR circuit with inductance  $10 \text{ H}$ , capacitance  $10 \mu\text{F}$ , resistance  $50\Omega$  is connected to an ac source of voltage,  $V = 200 \sin(100t)$  volt. If the resonant frequency of the LCR circuit is  $\nu_0$  and the frequency of the ac source is  $\nu$ , then  
 (1)  $\nu_0 = \nu = 50\text{Hz}$       (2)  $\nu_0 = \nu = \frac{50}{\pi}\text{Hz}$   
 (3)  $\nu_0 = \frac{50}{\pi}\text{Hz}, \nu = 50\text{Hz}$       (4)  $\nu = 100 \text{ Hz}; \nu_0 = \frac{100}{\pi}\text{Hz}$
46. Two pendulums of length  $121 \text{ cm}$  and  $100 \text{ cm}$  start vibrating in phase. At some instant, the two are at their mean position in the same phase, the minimum number of vibrations of the shorter pendulum after which the two are again in phase at the mean position is :  
 (1) 11      (2) 9      (3) 10      (4) 8
47. The volume occupied by the molecules contained in  $4.5 \text{ kg}$  water at STP, if the intermolecular forces vanish away is :  
 (1)  $5.6 \times 10^6 \text{ m}^3$       (2)  $5.6 \times 10^3 \text{ m}^3$       (3)  $5.6 \times 10^{-3} \text{ m}^3$       (4)  $5.6 \text{ m}^3$

48. Match List-I with List-II



	List-I		List-II
(a)	Gravitational constant (G)	(i)	$[L^2T^{-2}]$
(b)	Gravitational potential energy	(ii)	$[M^{-1}L^3T^{-2}]$
(c)	Gravitational potential	(iii)	$[LT^{-2}]$
(d)	Gravitational intensity	(iv)	$[ML^2T^{-2}]$

Choose the correct answer from the options given below :

- (1) (a) – (ii), (b)–(i), (c)–(iv), (d) – (iii)      (2) (a) – (ii), (b)–(iv), (c)–(i), (d) – (iii)  
 (3) (a) – (ii), (b)–(iv), (c)–(iii), (d) – (i)      (4) (a) – (iv), (b)–(ii), (c)–(i), (d) – (iii)
49. From Ampere's circuital law for a long straight wire of circular cross-section carrying a steady current, the variation of magnetic field in the inside and outside region of the wire :
- (1) uniform and remains constant for both the regions.  
 (2) a linearly increasing function of distance upto the boundary of the wire and then linearly decreasing for the outside region  
 (3) a linearly increasing function of distance  $r$  upto boundary of the wire and then decreasing one with  $1/r$  dependence for the outside region.  
 (4) a linearly decreasing function of distance upto the boundary of the wires and then a linearly increasing one for the outside region.
50. A nucleus of mass number 180 splits into two nuclei having mass number 125 and 64. The ratio of radius of two daughter nuclei respectively is :
- (1) 1 : 1      (2) 4 : 5      (3) 5 : 4      (4) 25 : 16



SECTION - A (CHEMISTRY)

51. Identify the incorrect statement from the following
- (1) Alkali metals react with water to form their hydroxides.
  - (2) The oxidation number of K in  $\text{KO}_2$  is + 4.
  - (3) Ionization enthalpy of alkali metals decreases from top to bottom in the group.
  - (4) Lithium is the strongest reducing agent among the alkali metals.
52. The IUPAC name of an element with atomic number 119 is:
- (1) ununennium
  - (2) unnilennium
  - (3) unununnium
  - (4) ununoctium
53. Which of the following is suitable to synthesize chlorobenzene?
- (1) Benzene,  $\text{Cl}_2$ , anhydrous  $\text{FeCl}_3$
  - (2) Phenol,  $\text{NaNO}_2$ ,  $\text{HCl}$ ,  $\text{CuCl}$
  - (3) ,  $\text{HCl}$
  - (4) ,  $\text{HCl}$ , Heating
54. Match List- I with List-II
- | List -I | List-II                                |
|---------|--|
| (a) Li  | (i) absorbent for carbon dioxide       |
| (b) Na  | (ii) electrochemical cells             |
| (c) KOH | (iii) coolant in fast breeder reactors |
| (d) Cs  | (iv) photoelectric cell                |
- Choose the correct answer from the options given below:
- (1) (a) - (iv), (b) - (i), (c) - (iii), (d) - (ii)
  - (2) (a) - (iii), (b) - (iv), (c) - (ii), (d). (i)
  - (3) (a) - (i), (b) - (iii), (c). (iv), (d) - (ii)
  - (4) (a)- (ii), (b) - (iii), (c) - (i), (d) - (iv)
55. Given below are two statements: one is labelled as Assertion (A) and the other is labelled as Reason (R).
- Assertion (A):**  $\text{ICl}$  is more reactive than  $\text{I}_2$ .
- Reason (R):**  $\text{I-Cl}$  bond is weaker than  $\text{I-I}$  bond.
- In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both (A) and (R) are correct and (R) is the correct explanation of (A)
  - (2) Both (A) and (R) are correct but (R) is not the correct explanation of (A).
  - (3) (A) is correct but (R) is not correct.
  - (4) (A) is not correct but (R) is correct.
56. Given below are two statements: one is labelled as Assertion (A) and the other is labelled as Reason (R).
- Assertion (A):** In a particular point defect, an ionic solid is electrically neutral, even if few of its cations are missing from its unit cells.
- Reason (R):** In an ionic solid, Frenkel defect arises due to dislocation of cation from its lattice site to interstitial site, maintaining overall electrical neutrality.
- In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both (A) and (R) are correct and (R) is the correct explanation of (A)
  - (2) Both (A) and (R) are correct but (R) is not the correct explanation of (A).
  - (3) (A) is correct but (R) is not correct.
  - (4) (A) is not correct but (R) is correct.

57. Given below are two statements:

**Statement I:**

The boiling points of aldehydes and ketones are higher than hydrocarbons of comparable molecular masses because of weak molecular association in aldehydes and ketones due to dipole - dipole interactions.

**Statement II:**

The boiling points of aldehydes and ketones are lower than the alcohols of similar molecular masses due to the absence of H-bonding.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct
- (2) Both Statement I and Statement II are incorrect.
- (3) Statement I is correct but Statement II is incorrect.
- (4) Statement I is incorrect but Statement II is correct.

58. Choose the correct statement:

- (1) Diamond and graphite have two-dimensional network.
- (2) Diamond is covalent and graphite is ionic.
- (3) Diamond is  $sp^3$  hybridized and graphite is  $sp^2$  hybridized.
- (4) Both diamond and graphite are used as dry lubricants.

59. Match List-I with List-II.

List -I	List-II
(Drug class)	(Drug molecule)
(a) Antacids	(i) Salvarsan
(b) Antihistamines	(ii) Morphine
(c) Analgesics	(iii) Cimetidine
(d) Antimicrobials	(iv) Seldane

Choose the correct answer from the options given below:

- (1) (a) - (iii), (b) - (ii), (c) - (iv), (d) - (i)
- (2) (a) - (iii), (b) - (iv), (c) - (ii), (d) - (i)
- (3) (a) - (i), (b) - (iv), (c) - (ii), (d) - (iii)
- (4) (a) - (iv), (b) - (iii), (c) - (i), (d) - (ii)

60. Match List -I with List -II.

List -I	List-II
(Products formed)	(Reaction of carbonyl compound with)
(a) Cyanohydrin	(i) $NH_2OH$
(b) Acetal	(ii) $RNH_2$
(c) Schiff's base	(iii) alcohol
(d) Oxime	(iv) $HCN$

Choose the correct answer from the options given below:

- (1) (a) - (iii), (b) - (iv), (c) - (ii), (d) - (i)
- (2) (a) - (ii), (b) - (iii), (c) - (iv), (d) - (i)
- (3) (a) - (i), (b) - (iii), (c) - (ii), (d) - (iv)
- (4) (a) - (iv), (b) - (iii), (c) - (ii), (d) - (i)

61. Given below are two statements:

**Statement I:**

Primary aliphatic amines react with  $\text{HNO}_2$  to give unstable diazonium salts.

**Statement II:**

Primary aromatic amines react with  $\text{HNO}_2$  to form diazonium salts which are stable even above 300 K.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct.
- (2) Both Statement I and Statement II are incorrect.
- (3) Statement I is correct but Statement II is incorrect.
- (4) Statement I is incorrect but Statement II is correct.

62. Given below are two statements:

**Statement I:**

In the coagulation of a negative sol, the flocculating power of the three given ions is in the order-  
 $\text{Al}^{3+} > \text{Ba}^{2+} > \text{Na}^+$

**Statement II:**

In the coagulation of a positive sol, the flocculating power of the three given salts is in the order-  
 $\text{NaCl} > \text{Na}_2\text{SO}_4 > \text{Na}_3\text{PO}_4$

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct.
- (2) Both Statement I and Statement II are incorrect.
- (3) Statement I is correct but Statement II is incorrect.
- (4) Statement I is incorrect but Statement II is correct.

63. Given below are two statements:

**Statement I:**

The boiling points of the following hydrides of group 16 elements increases in the order-  
 $\text{H}_2\text{O} < \text{H}_2\text{S} < \text{H}_2\text{Se} < \text{H}_2\text{Te}$

**Statement II:**

The boiling points of these hydrides increase with increase in molar mass.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct
- (2) Both Statement I and Statement II are incorrect
- (3) Statement I is correct but Statement II is incorrect
- (4) Statement I is incorrect but Statement II is correct

64. In one molal solution that contains 0.5 mole of a solute, there is:

- (1) 500 mL of solvent
- (2) 500 g of solvent
- (3) 100 mL of solvent
- (4) 1000 g of solvent

65. Which of the following statement is not correct about diborane?

- (1) There are two 3-centre-2-electron bonds.
- (2) The four terminal B-H bonds are two centre two electron bonds.
- (3) The four terminal Hydrogen atoms and the two Boron atoms lie in one plane.
- (4) Both the Boron atoms are  $\text{sp}^2$  hybridised.

66. Match List-I with List-II.

List -I(Hydrides)	List – II (Nature)
(a) $MgH_2$	(i) Electron precise
(b) $GeH_4$	(ii) Electron deficient
(c) $B_2H_6$	(iii) Electron rich
(d) HF	(iv) Ionic

Choose the correct answer from the options given below:

- (1) (a) - (iv), (b) - (i), (c) - (ii), (d) - (iii)      (2) (a) - (iii), (b) - (i), (c) - (ii), (d) - (iv)  
 (3) (a) - (i), (b) - (ii), (c) - (iv), (d) - (iii)      (4) (a) - (ii), (b) - (iii), (c) - (iv), (d) - (i)

67. The incorrect statement regarding chirality is:

- (1)  $S_N1$  reaction yields 1 : 1 mixture of both enantiomers.  
 (2) The product obtained by  $S_N2$  reaction of haloalkane having chirality at the reactive site shows inversion of configuration.  
 (3) Enantiomers are superimposable mirror images on each other.  
 (4) A racemic mixture shows zero optical rotation.

68. The pH of the solution containing 50 mL each of 0.10 M sodium acetate and 0.01 M acetic acid is: [Given  $pK_a$  of  $CH_3COOH = 4.57$ ]

- (1) 5.57      (2) 3.57      (3) 4.57      (4) 2.57

69. The IUPAC name of the complex -

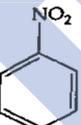

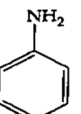
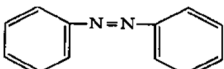
$[Ag(H_2O)_2][Ag(CN)_2]$  is:

- (1) dicyanidosilver(II) diaquaargentate(II)      (2) diaquasilver(II) dicyanidoargentate(II)  
 (3) dicyanidosilver (I) diaquaargentate (I)      (4) diaquasilver(I) dicyanidoargentate (I)

70. Which compound amongst the following is not an aromatic compound?

- (1)       (2)       (3)       (4) 

71. The Kjeldahl's method for the estimation of nitrogen can be used to estimate the amount of nitrogen in which one of the following compounds?

- (1)       (2)       (3)       (4) 

72. The incorrect statement regarding enzymes is:

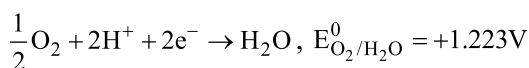
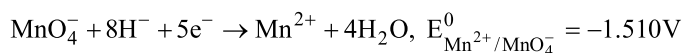
- (1) Enzymes are biocatalysts.  
 (2) Like chemical catalysts enzymes reduce the activation energy of bio processes.  
 (3) Enzymes are polysaccharides.  
 (4) Enzymes are very specific for a particular reaction and substrate.

73. Gadolinium has a low value of third ionization enthalpy because of:

- (1) small size      (2) high exchange enthalpy  
 (3) high electronegativity      (4) high basic character

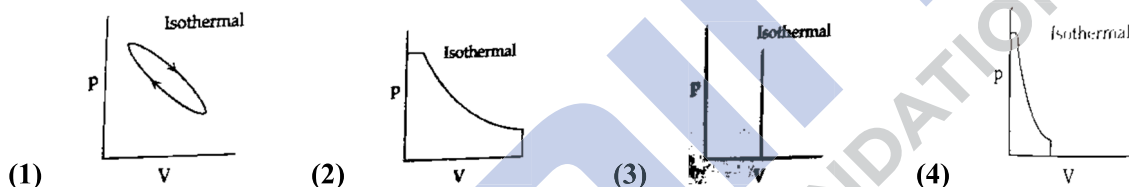
74. Which amongst the following is incorrect statement?
- (1) The bond orders of  $O_2^+$ ,  $O_2$ ,  $O_2^-$  and  $O_2^{2-}$  are 2.5, 2, 1.5 and 1 respectively.
  - (2)  $C_2$  molecule has four electrons in its two degenerate  $tt$  molecular orbitals.
  - (3)  $H_2^+$  ion has one electron.
  - (4)  $O_2^+$  ion is diamagnetic

75. Given below are half-cell reactions:



Will the permanganate ion,  $MnO_4^-$  liberate  $O_2$  from water in the presence of an acid?

- (1) Yes, because  $E_{cell}^0 = +0.287V$
  - (2) No, because  $E_{cell}^0 = -0.287V$
  - (3) Yes, because  $E_{cell}^0 = +2.733V$
  - (4) No, because  $E_{cell}^0 = -2.733V$
76. Which of the following p-V curve represents maximum work done?



77. Given below are two statements:

**Statement I:**

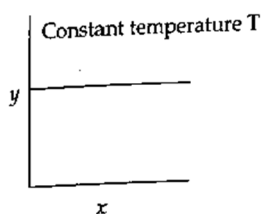
The acidic strength of monosubstituted nitrophenol is higher than phenol because of electron withdrawing nitro group.

**Statement II:**

*o*-nitrophenol, *m*-nitrophenol and *p*-nitrophenol will have same acidic strength as they have one nitro group attached to the phenolic ring.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct.
  - (2) Both Statement I and Statement II are incorrect.
  - (3) Statement I is correct but Statement II is incorrect.
  - (4) Statement I is incorrect but Statement II is correct.
78. The given graph is a representation of kinetics of a reaction.



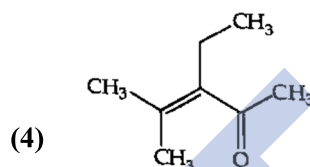
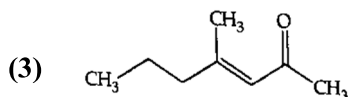
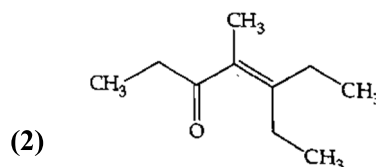
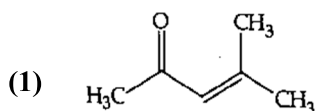
The y and x axes respectively are for zero and first order reactions, respectively are:

- (1) zero order ( $y = \text{concentration}$  and  $x = \text{time}$ ), first order ( $y = t_{1/2}$  and  $x = \text{concentration}$ )
- (2) zero order ( $y = \text{concentration}$  and  $x = \text{time}$ ), first order ( $y = \text{rate constant}$  and  $x = \text{concentration}$ )
- (3) zero order ( $y = \text{rate}$  and  $x = \text{concentration}$ ), first order ( $y = t_{1/2}$  and  $x = \text{concentration}$ )
- (4) zero order ( $y = \text{rate}$  and  $x = \text{concentration}$ ), first order ( $y = \text{rate}$  and  $y = t_{1/2}$ )

79. Identify the incorrect statement from the following.
- (1) All the five 5d orbitals are different in size when compared to the respective 4d orbitals.
  - (2) All the five 4d orbitals have shapes similar to the respective 3d orbitals.
  - (3) In an atom, all the five 3d orbitals are equal in energy in free state.
  - (4) The shapes of  $d_{xy}$ ,  $d_{yz}$  and  $d_{zx}$  orbitals are similar to each other; and  $d_{x^2-y^2}$  and  $d_{z^2}$  similar to each other.
80.  $\text{RMgX} + \text{CO}_2 \xrightarrow[\text{ether}]{\text{dry}} \text{Y} \xrightarrow{\text{H}_3\text{O}^+} \text{RCOOH}$   
What is Y in the above reaction?
- (1)  $\text{RCOO}^- \text{Mg}^+ \text{X}$
  - (2)  $\text{R}_3\text{CO}^- \text{Mg}^+ \text{X}$
  - (3)  $\text{RCOO}^- \text{X}^+$
  - (4)  $(\text{RCOO})_2 \text{Mg}$
81. Amongst the following which one will have maximum lone pair -lone pair' electron repulsions?
- (1)  $\text{ClF}_3$
  - (2)  $\text{IF}_5$
  - (3)  $\text{SF}_4$
  - (4)  $\text{XeF}_2$
82. Which one is not correct mathematical equation for ton s Law of partial pressure? Here p = total pressure of gaseous mixture.
- (1)  $p = p_1 + p_2 + p_3$
  - (2)  $p = n_1 \frac{RT}{V} + n_2 \frac{RT}{V} + n_3 \frac{RT}{V}$
  - (3)  $p_i = \chi_i p$ , where  $p_i$  = partial pressure of  $i^{\text{th}}$  gas
  - (4)  $p_i = \chi_i p_i^0$ , where  $\chi_i$  = mole fraction of  $i^{\text{th}}$  gas in gaseous mixture,  $p_i^0$  = pressure of  $i^{\text{th}}$  gas in pure state.
83. Which statement regarding polymers is not correct?
- (1) Elastomers have polymer chains held together by weak intermolecular forces.
  - (2) Fibers possess high tensile strength.
  - (3) Thermoplastic polymers are capable of repeatedly softening and hardening on heating and cooling respectively.
  - (4) Thermosetting polymers are reusable.
84. What mass of 95% pure  $\text{CaCO}_3$  will be require neutralize 50 mL of 0.5 M HCl solution according to the following reaction?
- $$\text{CaCO}_{3(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + \text{CO}_{2(g)} + 2\text{H}_2\text{O}_{(l)}$$
- [Calculate up to second place of decimal point]
- (1) 1.25 g
  - (2) 1.32 g
  - (3) 3.65 g
  - (4) 9.50 g
85. At 298 K, the standard electrode potentials of  $\text{Cu}^{2+} / \text{Cu}$ ,  $\text{Zn}^{2+} / \text{Zn}$ ,  $\text{Fe}^{2+} / \text{Fe}$  and  $\text{Ag}^+ / \text{Ag}$  are 0.34V, -0.76V, -0.44 V and 0.80 V, respectively. On the basis of standard electrode potential, predict which of the following reaction cannot occur?
- (1)  $\text{CuSO}_4(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu}(\text{s})$
  - (2)  $\text{CuSO}_4(\text{aq}) + \text{Fe}(\text{s}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu}(\text{s})$
  - (3)  $\text{FeSO}_4(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Fe}(\text{s})$
  - (4)  $2\text{CuSO}_4(\text{aq}) + 2\text{Ag}(\text{s}) \rightarrow 2\text{Cu}(\text{s}) + \text{Ag}_2\text{SO}_4(\text{aq})$

SECTION - B (CHEMISTRY)

86. Which one of the following is not formed where acetone reacts with 2-pentanone in the presence of dilute NaOH followed by heating?



87. For a first order reaction  $A \rightarrow \text{Products}$ , initial concentration of A is 0.1 M, which becomes 0.001 after 5 minutes. Rate constant for the reaction  $\text{min}^{-1}$  is:

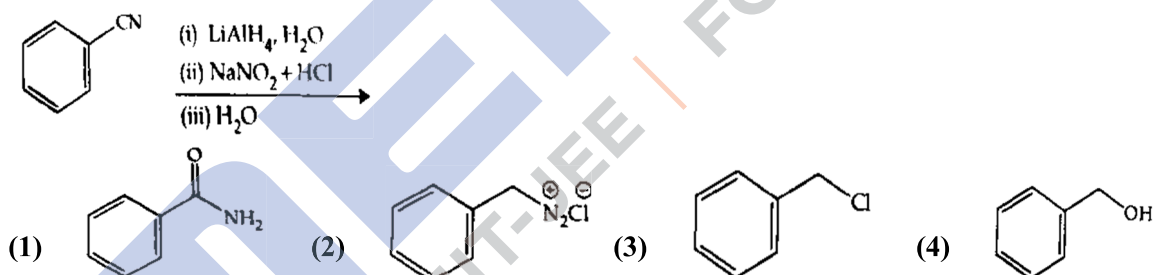
- (1) 1.3818                      (2) 0.9212                      (3) 0.4606                      (4) 0.2303

88.  $3\text{O}_2(\text{g}) \rightleftharpoons 2\text{O}_3(\text{g})$

For the above reaction at 298 K,  $K_C$  is found to be  $3.0 \times 10^{-59}$ . If the concentration of  $\text{O}_2$  at equilibrium is 0.040 M then the concentration of  $\text{O}_3$  in M is:

- (1)  $4.38 \times 10^{-32}$                       (2)  $1.9 \times 10^{-63}$                       (3)  $2.4 \times 10^{31}$                       (4)  $1.2 \times 10^{21}$

89. The product formed from the following reaction sequence is:



90. Match List -I with List -I

List -I (Ores)

List-II (Composition)

- |               |  |
|---------------|--|
| (a) Hematite  | (i) $\text{Fe}_3\text{O}_4$                            |
| (b) Magnetite | (ii) $\text{ZnCO}_3$                                   |
| (c) Calamine  | (iii) $\text{Fe}_2\text{O}_3$                          |
| (d) Kaolinite | (iv) $[\text{Al}_2(\text{OH})_4\text{Si}_2\text{O}_5]$ |

Choose the correct answer from the options given below:

- (1) (a) - (i), (b) - (ii), (c) - (iii), (d) - (iv)                      (2) (a)-(iii), (b) - (i), (c) - (ii), (d)-(iv)  
 (3) (a) - (iii), (b) - (i), (c) - (iv), (d) - (ii)                      (4) (a) - (i), (b) - (iii), (c) - (ii), (d) - (iv)

91. Given below are two statements:

**Statement I:** In Lucas test, primary, secondary and tertiary alcohols are distinguished on the basis of their reactivity with  $\text{conc. HCl} + \text{ZnCl}_2$ , known as Lucas Reagent.

**Statement II:** Primary alcohols are most reactive and immediately produce turbidity at room temperature on reaction with Lucas Reagent

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct.  
 (2) Both Statement I and Statement II are incorrect.  
 (3) Statement I is correct, but Statement II is incorrect.  
 (4) Statement I is incorrect, but Statement II is correct.
92. Find the emf of the cell in which the following reaction takes place at 298 K  

$$\text{Ni(s)} + 2\text{Ag}^+(0.001\text{M}) \rightarrow \text{Ni}^{2+}(0.001\text{M}) + 2\text{Ag(s)}$$
 (Given that  $E_{\text{cell}}^0 = 10.5\text{V}$ ,  $\frac{2.303RT}{F} = 0.059$  at 298 K)  
 (1) 1.0385 V (2) 1.385 V (3) 0.9615 V (4) 1.05 V
93. Compound X on reaction with  $\text{O}_3$  followed by  $\text{Zn}/\text{H}_2\text{O}$  gives formaldehyde and 2-methyl propanal as products. The compound X is:  
 (1) 3-Methylbut-1-ene (2) 2-Methylbut-1-ene (3) 2-Methylbut-2-ene (4) Pent-2-ene
94. In the neutral or faintly alkaline medium,  $\text{KMnO}_4$  oxidizes iodide into iodate. The change in oxidation state of manganese in this reaction is from:  
 (1) +7 to +4 (2) +6 to +4 (3) +7 to +3 (4) +6 to +5
95. Copper crystallizes in fcc unit cell with cell edge length of  $3.608 \times 10^{-8} \text{ cm}$ . The density of copper is  $8.92 \text{ g cm}^{-3}$ . Calculate the atomic mass of copper.  
 (1) 63.1 u (2) 31.55 u (3) 60 u (4) 65 u
96. A 10.0 L flask contains 64 g of oxygen at  $27^\circ\text{C}$ . (Assume  $\text{O}_2$  gas is behaving ideally). The pressure inside the flask in bar is (Given  $R = 0.0831 \text{ L bar K}^{-1} \text{ mol}^{-1}$ )  
 (1) 2.5 (2) 408.6 (3) 40.8 (4) 4.0
97. If radius of second Bohr orbit of the  $\text{He}^+$  ion is 105.8 pm, what is the radius of third Bohr orbit of  $\text{Li}^{2+}$  ion?  
 (1) 158.7 pm (2) 15.87 pm (3) 1.587 pm (4) 158.7 Å
98. The order of energy absorbed which is responsible for the color of complexes.  
 (A)  $[\text{Ni}(\text{H}_2\text{O})_2(\text{en})_2]^{2+}$  (B)  $[\text{Ni}(\text{H}_2\text{O})_4(\text{en})]^{2+}$  (C)  $[\text{Ni}(\text{en})_3]^{2+}$   
 (1) (A) > (B) > (C) (2) (C) > (B) > (A) (3) (C) > (A) > (B) (4) (B) > (A) > (C)
99. The correct IUPAC name of the following compound is:
- 
- (1) 1-bromo-5-chloro-4-methylhexan-3-ol (2) 6-bromo-2-chloro-4-methylhexan-4-ol  
 (3) 1-bromo-4-methyl-5-chlorohexan-3-ol (4) 6-bromo-4-methyl-2-chlorohexan-4-ol
100. The pollution due to oxides of sulphur get enhanced due to the presence of:  
 (a) particulate matter (b) ozone (c) hydrocarbons (d) hydrogen peroxide  
 Choose the most appropriate answer from the options given below:  
 (1) (a), (d) only (2) (a), (b), (d) only (3) (b), (c) (d) only (4) (a), (c), (d) only



## SECTION - A (BIOLOGY : BOTANY)

101. The process of translation of mRNA to proteins begins as soon as:
- (1) The small subunit of ribosome encounters mRNA
  - (2) The larger subunit of ribosome encounters mRNA
  - (3) Both the subunits join together to bind with mRNA
  - (4) The tRNA is activated and the larger subunit of ribosome encounters mRNA
102. The device which can remove particulate matter present in the exhaust from a thermal power plant is:
- (1) STP
  - (2) Incinerator
  - (3) Electrostatic Precipitator
  - (4) Catalytic Converter
103. Which of the following is incorrectly matched?
- (1) *Ectocarpus* - Fucoxanthin.
  - (2) *Ulothrix* - Mannitol
  - (3) *Porphyra* - Floridian Starch
  - (4) *Volvox* - Starch
104. Hydrocolloid carrageen is obtained from:
- (1) Chlorophyceae and Phaeophyceae
  - (2) Phaeophyceae and Rhodophyceae
  - (3) Rhodophyceae only
  - (4) Phaeophyceae only
105. Which one of the following statements cannot be connected to Predation?
- (1) It helps in maintaining species diversity in a community
  - (2) It might lead to extinction of a species
  - (3) Both the interacting species are negatively impacted
  - (4) It is necessitated by nature to maintain the ecological balance
106. Given below are two statements:  
**Statement I:** The primary CO<sub>2</sub> acceptor in C<sub>4</sub> plants is phosphoenolpyruvate and is found in the mesophyll cells.  
**Statement II:** Mesophyll cells of C<sub>4</sub> plants lack RuBisCo enzyme.  
In the light of the above statements, choose the correct answer from the options given below :
- (1) Both Statement I and Statement II are correct
  - (2) Both Statement I and Statement II are incorrect
  - (3) Statement I is correct but Statement II is incorrect
  - (4) Statement I is Incorrect but Statement II is correct
107. Which one of the following produces nitrogen fixing nodules on the roots of *Alnus* ?
- (1) *Rhizobium*
  - (2) *Frankia*
  - (3) *Rhodospirillum*
  - (4) *Beijernickia*
108. DNA polymorphism forms the basis of:
- (1) Genetic mapping
  - (2) DNA finger printing
  - (3) Both genetic mapping and DNA finger printing
  - (4) Translation
109. Which one of the following plants does not show plasticity?
- (1) Cotton
  - (2) Coriander
  - (3) Buttercup
  - (4) Maize
110. What is the net gain of ATP when each molecule of glucose is converted to two molecules of pyruvic acid
- (1) Four
  - (2) Six
  - (3) Two
  - (4) Eight

111. In old trees the greater part of secondary xylem is dark brown and resistant to insect attack due to :
- secretion of secondary metabolites and their deposition in the lumen of vessels.
  - deposition of organic compounds like tannins and resins in the central layers of stem.
  - deposition of suberin and aromatic substances in the outer layer of stem
  - deposition of tannins, gum, resin and aromatic substances in the peripheral layers of stem.
  - presence of parenchyma cells, functionally active xylem elements and essential oils.
- Choose the correct answer from the options given below:
- (1) (a) and (b) Only    (2) (c) and (d) Only    (3) (d) and (e) Only    (4) (b) and (d) Only
112. The flowers are Zygomorphic in:
- Mustard
  - Gulmohar
  - Cassia
  - Datura
  - Chilly
- Choose the correct answer from the options given below:
- (1) (a), (b), (c) Only    (2) (b), (c) Only    (3) (d), (e) Only    (4) (c), (d), (e) Only
113. What amount of energy is released from glucose during lactic acid fermentation?
- Approximately 15%
  - More than 18%
  - About 10%
  - Less than 7%
114. The gaseous plant growth regulator is used in plant to:
- speed up the malting process
  - promote root growth and root hair formation to increase the absorption surface
  - help overcome apical dominance
  - kill dicotyledonous weeds in the fields
115. Identify the incorrect statement related to Pollination:
- Pollination by water is quite rare in flowering plants
  - Pollination by wind is more common amongst abiotic pollination
  - Flowers produce foul odours to attract flies and beetles to get pollinated
  - Moths and butterflies are the most dominant pollinating agents among insects
116. Habitat loss and fragmentation, over exploitation, alien species invasion and co-extinction are causes for:
- Population explosion
  - Competition
  - Biodiversity loss
  - Natality
117. The appearance of recombination nodules on homologous chromosomes during meiosis characterizes:
- Synaptonemal complex
  - Bivalent
  - Sites at which crossing over occurs
  - Terminalization
118. Production of Cucumber has increased manifold in recent years. Application of which of the following phytohormones has resulted in this increased yield as the hormone is known to produce female flowers in the plants:
- ABA
  - Gibberellin
  - Ethylene
  - Cytokinin
119. Which of the following is not a method of *ex situ* conservation?
- In vitro* fertilization
  - National Parks
  - Micropropagation
  - Cryopreservation
120. Which one of the following statement is not true regarding gel electrophoresis technique?
- The process of extraction of separated DNA strands from gel is called elution.
  - The separated DNA fragments are stained by using ethidium bromide.
  - The presence of chromogenic substrate given blue coloured DNA bands on the gel.
  - Bright orange coloured bands of DNA can be observed in the gel when exposed to UV light.

121. Read the following statements and choose the set of correct statements:

- (a) Euchromatin is loosely packed chromatin
- (b) Heterochromatin is transcriptionally active
- (c) Histone octamer is wrapped by negatively charged DNA in nucleosome
- (d) Histones are rich in lysine and arginine
- (e) A typical nucleosome contains 400 bp of DNA helix

Choose the correct answer from the options given below:

- (1) (a), (d), (e) only
- (2) (a), (c) (d) only
- (3) (b), (e) Only
- (4) (a), (c), (e) Only

122. Match List - I with List – II

	List-I		List-II
(a)	Manganese	(i)	Activates the enzyme catalase
(b)	Magnesium	(ii)	Required for pollen germination
(c)	Boron	(iii)	Activates enzymes of respiration
(d)	Iron	(iv)	Functions in splitting of water during photosynthesis

Choose the correct answer from the options given below :

- (1) (a)-(iii), (b), (iv), (c) - (i), (d)-(ii)
- (2) (a)-(iv), (b), (iii), (c) - (ii), (d)-(i)
- (3) (a)-(iv), (b), (i), (c) - (ii), (d)-(iii)
- (4) (a)-(iii), (b), (i), (c) - (ii), (d)-(iv)

123. Given below are two statements : one is labelled as Assertion (A) and the other is labelled as Reason (R)  
**Assertion (A):** Polymerase chain reaction is used in DNA amplification.

**Reason (R):** The ampicillin resistant gene is used as a selectable marker to check transformation.

In the light of the above statements, choose the correct answer from the options given below :

- (1) Both (A) and (R) are correct and (R) is the correct explanation of (A)
- (2) Both (A) and (R) are correct but (R) is not the correct explanation of (A)
- (3) (A) is correct but (R) is not correct
- (4) (A) is not correct but (R) is correct

124. Which one of the following never occurs during mitotic cell division?

- (1) Spindle fibres attach to kinetochores of chromosomes
- (2) Movement of centrioles towards opposite poles
- (3) Pairing of homologous chromosomes
- (4) Coiling and condensation of the chromatids

125. Which of the following is not observed during apoplastic pathway?

- (1) Movement of water occurs through intercellular spaces and wall of the cells.
- (2) The movement does not involve crossing of cell membrane
- (3) The movement is aided by cytoplasmic streaming
- (4) Apoplast is continuous and does not provide any barrier to water movement.

126. Given below are two statements:

**Statement I :** Cleistogamous flowers are invariably autogamous

**Statement II :** Cleistogamy is disadvantageous as there is no chance for cross pollination.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Both Statement I and Statement II are correct
- (2) Both Statement I and Statement II are incorrect
- (3) Statement I is correct but Statement II is incorrect
- (4) Statement I is incorrect but Statement II is correct

127. "Girdling Experiment" was performed by Plant Physiologists to identify the plant tissue through which  
 (1) water is transported (2) food is transported  
 (3) for both water and food transportation (4) osmosis is observed
128. XO type of sex determination can be found in:  
 (1) *Drosophila* (2) Birds (3) Grasshoppers (4) Monkeys
129. Read the following statements about the vascular bundles:  
 (a) In roots, xylem and phloem in a vascular bundle are arranged in an alternate manner along the different radii.  
 (b) Conjoint closed vascular bundles do not possess cambium.  
 (c) In open vascular bundles, cambium is present in between xylem and phloem  
 (d) The vascular bundles of dicotyledonous stem possess endarch protoxylem  
 (e) In monocotyledonous root, usually there are more than six xylem bundles present.  
 Choose the correct answer from the options given below:  
 (1) (a), (b) and (d) Only (2) (b), (c), (d) and (e) Only  
 (3) (a), (b), (c) and (d) Only (4) (a), (c), (d) and (e) Only
130. Which one of the following plants shows vexillary aestivation and diadelphous stamens ?  
 (1) *Colchicum autumnale* (2) *Pisum sativum*  
 (3) *Allium cepa* (4) *Solanum nigrum*
131. Given below are two statements  
**Statement I:** Decomposition is a process in which the detritus is degraded into simpler substances by microbes.  
**Statement II:** Decomposition is faster if the detritus is rich in lignin and chitin  
 In the light of the above statements, choose the correct answer from the options given below :  
 (1) Both Statement I and Statement II are correct  
 (2) Both Statement I and Statement II are incorrect  
 (3) Statement I is correct but Statement II is incorrect  
 (4) Statement I is incorrect but Statement II is correct
132. Identify the correct set of statements:  
 (a) The leaflets are modified into pointed hard thorns in *Citrus* and *Bougainvillea*.  
 (b) Axillary buds form slender and spirally collet tendrils in cucumber and pumpkin  
 (c) Stern is flattened and fleshy in *Opuntia* and modified to perform the function of leaves  
 (d) *Rhizophora* shows vertically upward growing roots that help to get oxygen for respiration  
 (e) Subaerially growing stems in grasses and strawberry help in vegetative propagation  
 Choose the correct answer from the options given below:  
 (1) (b) and (c) Only (2) (a) and (a) Only  
 (3) (1i), (c), (d) and (e) Only (4) (a), (b), (d) and (e) Only
133. Exoskeleton of arthropods is composed of:  
 (1) Cutin (2) Cellulose (3) Chitin (4) Glucosamine
134. Which one of the following is not true regarding the release of energy during ATP synthesis through chemiosmosis? It involves:  
 (1) Breakdown of proton gradient  
 (2) Breakdown of electron gradient  
 (3) Movement of protons across the membrane to the stroma  
 (4) Reduction of NADP to NADPH<sub>2</sub> on the stroma side of the membrane

135. Given below are two statements

**Statement I:** Mendel studied seven pairs of contrasting traits in pea plants and proposed the Laws of Inheritance

**Statement II:** Seven characters examined by Mendel in his experiment on pea plants were seed shape and colour, flower colour, pod shape and colour, flower position and stem height

In the light of the above statements, choose the correct answer from the options given below :

- (1) Both Statement I and Statement II are correct
- (2) Both Statement I and Statement II are incorrect
- (3) Statement I is correct but Statement II is incorrect
- (4) Statement I is incorrect but Statement II is correct

**SECTION - B (BIOLOGY : BOTANY)**

136. Match the plant with the kind of life cycle it exhibits:

	List - I		List-II
(a)	<i>Spirogyra</i>	(i)	Dominant diploid sporophyte vascular plant, with highly reduced male or female gametophyte
(b)	Fern	(ii)	Dominant haploid free-living gametophyte
(c)	<i>Funaria</i>	(iii)	Dominant diploid sporophyte alternating with reduced gametophyte called prothallus
(d)	<i>Cycas</i>	(iv)	Dominant haploid leafy gametophyte alternating with partially dependent multicellular sporophyte

Choose the correct answer from the options given below:

- (1) (a)-(iv), (b)-(i), (c)-(ii), (d)-(iii)
- (2) (a)-(ii), (b)-(iii), (c)-(iv), (d)-(i)
- (3) (a)-(iii), (b)-(iv), (c)-(i), (d)-(ii)
- (4) (a)-(ii), (b)-(iv), (c)-(i), (d)-(iii)

137. The anatomy of springwood shows some peculiar features. Identify the correct set of statements about springwood.

- (a) It is also called as the earlywood
- (b) In spring season cambium produces xylem elements with narrow vessels
- (c) It is lighter in colour
- (d) The springwood along with autumnwood shows alternate concentric rings forming annual rings
- (e) It has lower density

Choose the correct answer from the options given below:

- (1) (a), (b), (d) and (e) Only
- (2) (a), (c), (d) and (e) Only
- (3) (a), (b) and (d) Only
- (4) (c), (d) and (e) Only

138. In the following palindromic base sequences of DNA, which one can be cut easily by particular restriction enzyme?

- (1) 5' G A C A C T 3' ; 3' C T A T G A 5'
- (2) 5' G A A T T C 3' ; 3' C T T A A G 5'
- (3) 5' G T A T T C 3' ; 3' C A T A A G 5'
- (4) 5' G T A T T C 3' ; 3' C A T A A G 5'

139. While explaining interspecific interaction of population, (+) sign is assigned for beneficial interaction, (–) sign is assigned for detrimental interaction and (0) for neutral interaction. Which of the following interactions can be assigned (+) for one species and (–) for another species involved in the interaction?

- (1) Predation
- (2) Amensalism
- (3) Commensalism
- (4) Competition

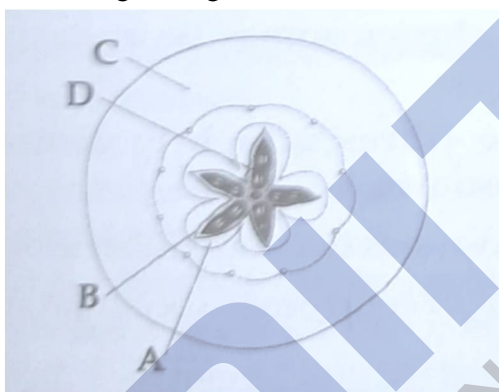
140. Given below are two statements : one is labelled as Assertion (A) and the other is labelled as Reason (R).  
**Assertion (A):** Mendel's law of Independent assortment does not hold good for the genes that are located closely on the same chromosome.

**Reason (R):** Closely located genes assort independently.

In the light of the above statements, choose the correct answer from the options given below :

- (1) Both (A) and (R) are correct and (R) is the correct explanation of (A)  
 (2) Both (A) and (R) are correct but (R) is not the correct explanation of (A)  
 (3) (A) is correct but (R) is not correct  
 (4) (A) is not correct but (R) is correct

141. Which part of the fruit, labelled in the given figure makes it a false fruit ?



- (1) A → Mesocarp      (2) B → Endocarp      (3) C → Thalamus      (4) D → Seed

142. Match List - I with List - II

	List-I		List-II
(a)	Metacentric chromosome	(i)	Centromere situated close to the end forming one extremely short and one very long arms.
(b)	Acrocentric chromosome	(ii)	Centromere at the terminal end
(c)	Sub-metacentric	(iii)	Centromere in the middle forming two equal arms of chromosomes
(d)	Telocentric chromosome	(iv)	Centromere slightly away from the middle forming one shorter arm and one longer arm

Choose the correct answer from the options given below:

- (1) (a) - (iii), (b) - (i), (c) - (iv), (d) - (ii)      (2) (a) - (i), (b) - (iii), (c) (ii), (d) - (iv)  
 (3) (a) - (ii), (b) - (iii), (c) - (iv), (d) - (i)      (4) (a) - (i), (b) - (ii), (c) - (iii), (d) - (iv)

143. Addition of more solutes in a given solution will:

- (1) raise its water potential      (2) lower its water potential  
 (3) make its water potential zero      (4) not affect the water potential at all

144. Which one of the following will accelerate phosphorus cycle?

- (1) Burning of fossil fuels      (2) Volcanic activity  
 (3) Weathering of rocks      (4) Rain fall and storms

145. Which of the following occurs due to the presence of autosomal dominant trait?

- (1) Sickle cell anaemia      (2) Myotonic dystrophy  
 (3) Haemophilia      (4) Thalassemia

146. Read the following statements on lipids and find out correct set of statements:
- (a) Lecithin found in the plasma membrane is a glycolipid.
  - (b) Saturated fatty acids possess one or more  $C=C$  bonds.
  - (c) Gingely oil has lower melting point, hence remains as oil in winter.
  - (d) Lipids are generally insoluble in water but soluble in some organic solvents
  - (e) When fatty acid is esterified with glycerol, monoglycerides are formed.
- Choose the correct answer from the options given below:
- (1) (a), (b) and (c) only
  - (2) (a), (d) and (e) only
  - (3) (c), (d) and (e) only
  - (4) (a), (b) and (d) only
147. What is the role of large bundle sheath cells found around the vascular bundles in  $C_4$  plants?
- (1) To provide the site for photorespiratory pathway
  - (2) To increase the number of chloroplast for the operation of Calvin cycle
  - (3) To enable the plant to tolerate high temperature
  - (4) To protect the vascular tissue from high light intensity.
148. The entire fleet of buses in Delhi were converted to CNG from diesel. In reference to this, which one of the following statements is false?
- (1) CNG burns more efficiently than diesel
  - (2) The same diesel engine is used in CNG buses making the cost of conversion low.
  - (3) It is cheaper than diesel
  - (4) It can not be adulterated like diesel
149. Transposons can be used during which one of the following?
- (1) Polymerase Chain Reaction
  - (2) Gene silencing
  - (3) Autoradiography
  - (4) Gene sequencing
150. If a geneticist uses the blind approach for sequencing the whole genome of an organism, followed by assignment of function to different segments, the methodology adopted by him is called as:
- (1) Sequence annotation
  - (2) Gene mapping
  - (3) Expressed sequence tags
  - (4) Bioinformatics

## SECTION - A (BIOLOGY : ZOOLOGY)

151. Given below are two statements: one is labelled as Assertion (A) and the other is labelled as Reason (R).  
Assertion (A): Osteoporosis is characterized by decreased bone mass and increases chances of fractures.  
Reason (R): Common cause of osteoporosis is increased levels of estrogen. In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both (A) and (R) are correct and (R) is the correct explanation (A)
  - (2) Both (A) and (R) are correct but (R) is not the correct explanation of (A)
  - (3) (A) is correct but (R) is not correct
  - (4) (A) is not correct but (R) is correct
152. A dehydration reaction links two glucose molecules to produce maltose. If the formula for glucose is  $C_6H_{12}O_6$  then what is the formula for maltose?
- (1)  $C_{12}H_{20}O_{10}$
  - (2)  $C_{12}H_{24}O_{12}$
  - (3)  $C_{12}H_{22}O_{11}$
  - (4)  $C_{12}H_{24}O_{11}$
153. In which of the following animals, digestive tract has additional chambers like crop and gizzard?
- (1) *Corvus, Columba, Chameleon*
  - (2) *Bufo, Balaenoptera, Bangarus*
  - (3) *Catla, Columba, Crocodilus*
  - (4) *Pavo, Psittacula, Corvus*
154. Select the incorrect statement with reference to mitosis:
- (1) All the chromosomes lie at the equator at metaphase
  - (2) Spindle fibres attach to centromere of chromosomes
  - (3) Chromosomes decondense at telophase
  - (4) Splitting of centromere occurs at anaphase.
155. Which of the following statements with respect to Endoplasmic Reticulum is incorrect?
- (1) RER has ribosomes attached to ER
  - (2) SER is devoid of ribosomes
  - (3) In prokaryotes only RER are present
  - (4) SER are the sites for lipid synthesis
156. Regarding Meiosis, which of the statements is incorrect?
- (1) There are two stages in Meiosis, Meiosis-I and II
  - (2) DNA replication occurs in S phase of Meiosis-I
  - (3) Pairing of homologous chromosomes and recombination occurs in Meiosis-I
  - (4) Four haploid cells are formed at the end of Meiosis-II
157. Breeding crops with higher levels of vitamins and minerals or higher proteins and healthier fats is called
- (1) Bio-magnification
  - (2) Bio-remediation
  - (3) Bio-fortification
  - (4) Bio-accumulation
158. Tegmina in cockroach, arises from:
- (1) Prothorax
  - (2) Mesothorax
  - (3) Metathorax
  - (4) Prothorax and Mesothorax
159. Given below are two statements:  
Statement I: Fatty acids and glycerols cannot be absorbed into the blood.  
Statement II: Specialized lymphatic capillaries called lacterals carry chylomicrons into lymphatic vessels and ultimately into the blood. In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both statement I and Statement II are correct
  - (2) Both statement I and Statement II are incorrect
  - (3) Statement I is correct but Statement II is incorrect



- (4) Statement I is incorrect but Statement II is correct
160. Given below are two statements:  
Statement I: The release of sperms into the seminiferous tubules is called spermiation.  
Statement II: Spermiogenesis is the process of formation of sperms from spermatogonia.  
In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both statement I and Statement II are correct  
(2) Both statement I and Statement II are incorrect  
(3) Statement I is correct but Statement II is incorrect  
(4) Statement I is incorrect but Statement II is correct
161. *In-situ* conservation refers to:
- (1) Protect and conserve the whole ecosystem      (2) Conserve only high risk species  
(3) Conserve only endangered species              (4) Conserve only extinct species
162. Given below are two statements:  
**Statement I:** Mycoplasma can pass through less than 1 micron filter size.  
**Statement II:** Mycoplasma are bacteria with cell wall  
In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both statement I and Statement II are correct  
(2) Both statement I and Statement II are incorrect  
(3) Statement I is correct but Statement II is incorrect  
(4) Statement I is incorrect but Statement II is correct
163. Which of the following is a correct match for disease and its symptoms?
- (1) Arthritis – Inflamed joints  
(2) Tetany – high  $\text{Ca}^{2+}$  level causing rapid spasms  
(3) Myasthenia gravis – Genetic disorder resulting in weakening and paralysis of skeletal muscle  
(4) Muscular dystrophy – An auto immune disorder causing progressive degeneration of skeletal muscle
164. Given below are two statements:  
Statement I: Autoimmune disorder is a condition where body defense mechanism recognizes its own cells as foreign bodies.  
Statement II: Rheumatoid arthritis is a condition where body does not attack self cells.  
In the light of the above statements, choose the most appropriate answer from the options given below:
- (1) Both statement I and Statement II are correct  
(2) Both statement I and Statement II are incorrect  
(3) Statement I is correct but Statement II is incorrect  
(4) Statement I is incorrect but Statement II is correct
165. In an *E.coli* strain *i* gene gets mutated and its product can not bind the inducer molecule. If growth medium is provided with lactose, what will be the outcome?
- (1) Only *z* gene will get transcribed                      (2) *z*, *y*, *a* genes will transcribed  
(3) *z*, *y* *a* genes will not be translated              (4) RNA polymerase will bind the promoter region
166. Which of the following statements are true for spermatogenesis but do not hold true for Oogenesis?
- (a) It results in the formation of haploid gametes  
(b) Differentiation of gamete occurs after the completion of meiosis  
(c) Meiosis occurs continuously in a mitotically dividing stem cell population  
(d) It is controlled by the Luteinising hormone (LH) and Follicle Stimulating Hormone (FSH) secreted

by the anterior pituitary

(e) It is initiated at puberty

Choose the most appropriate answer from the options given below:

- (1) (c) and (e) only                      (2) (b) and (c) only  
(3) (b), (d) and (e) only                (4) (b), (c) and (e) only

167. Under normal physiological conditions in human being every 100 ml of oxygenated blood can deliver \_\_\_\_\_ ml of O<sub>2</sub> to the tissues.  
(1) 2 ml                      (2) 5 ml                      (3) 4 ml                      (4) 10 ml
168. Nitrogenous waste is excreted in the form of pellet or paste by:  
(1) Ornithorhynchus    (2) Salamandra            (3) Hippocampus            (4) Pavo
169. Which of the following functions is not performed by secretions from salivary glands?  
(1) Control bacterial population in mouth            (2) Digestion of complex carbohydrates  
(3) Lubrication of oral cavity                            (4) Digestion of disaccharides
170. Natural selection where more individuals acquire specific character value other than the mean character value, leads to:  
(1) Stabilising change    (2) Directional change    (3) Disruptive change    (4) Random change
171. If the length of a DNA molecule is 1.1 metres, what will be the approximate number of base pairs?  
(1)  $3.3 \times 10^9$  bp            (2)  $6.6 \times 10^9$  bp            (3)  $3.3 \times 10^6$  bp            (4)  $6.6 \times 10^6$  bp
172. Which of the following is not a connective tissue?  
(1) Blood                      (2) Adipose tissue            (3) Cartilage                      (4) Neuroglia
173. Given below are two statements: one is labelled as Assertion (A) and the other is labelled as Reason (R).  
Assertion (A): All vertebrates are chordates but all chordates are not vertebrates.  
Reason (R): Notochord is replaced by vertebral column in the adult vertebrates  
In the light of the above statements, choose the most appropriate answer from the option given below:  
(1) Both (A) and (R) are correct and (R) is the correct explanation (A)  
(2) Both (A) and (R) are correct but (R) is not the correct explanation of (A)  
(3) (A) is correct but (R) is not correct  
(4) (A) is not correct but (R) is correct
174. In the taxonomic categories which hierarchical arrangement in ascending is correct in case of animals?  
(1) Kingdom, Phylum, Class, Order, Family, Genus, Species  
(2) Kingdom, Class, Phylum, Family, Order, Genus, Species  
(3) Kingdom, Order, Class, Phylum, Family, Genus, Species  
(4) Kingdom, Order, Phylum, Class, Family, Genus, Species
175. Identify the microorganism which is responsible for the production of an immunosuppressive molecule cyclosporin A:  
(1) *Trichoderma polysporum*                      (2) *Clostridium butylicum*  
(3) *Aspergillus niger*                                      (4) *Streptococcus cerevisiae*
176. If '8' Drosophila in a laboratory population of '80' died during a week, the death rate in the population is \_\_\_\_\_ individuals per Drosophila per week.

- (1) 0.1                      (2) 10                      (3) 1.0                      (4) zero

177. Given below are two statements:

Statement I: The coagulum is formed of network of threads called thrombins.

Statement II: Spleen is the graveyard of erythrocytes

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both statement I and Statement II are correct  
 (2) Both statement I and Statement II are incorrect  
 (3) Statement I is correct but Statement II is incorrect  
 (4) Statement I is incorrect but Statement II is correct

178. Which of the following is present between the adjacent bones of the vertebral column?

- (1) Intercalated discs    (2) Cartilage              (3) Areolar tissue        (4) Smooth muscle

179. Which of the following is not the function of conducting part of respiratory system?

- (1) It clears inhaled air from foreign particles  
 (2) Inhaled air is humidified  
 (3) Temperature of inhaled air is brought to body temperature  
 (4) Provides surface for diffusion of O<sub>2</sub> and CO<sub>2</sub>

180. Lippe's loop is a type of contraceptive used as:

- (1) Cervical barrier    (2) Vault barrier  
 (3) Non-Medicated IUD    (4) Copper releasing IUD

181. In gene therapy of Adenosine Deaminase (ADA) deficiency, the patient requires periodic infusion of genetically engineered lymphocytes because:

- (1) Retroviral vector is introduced into these lymphocytes  
 (2) Gene isolated from marrow cells producing ADA is introduced into cells at embryonic stages  
 (3) Lymphocytes from patient's blood are grown in culture, outside the body.  
 (4) Genetically engineered lymphocytes are not immortal cells

182. Detritivores breakdown detritus into smaller particles. This process is called:

- (1) Catabolism              (2) Fragmentation        (3) Humification              (4) Decomposition

183. Given below are two statements:

Statement I: Restriction endonucleases recognize specific sequence to cut DNA known as palindromic nucleotide sequence

Statement II: Restriction endonucleases cut DNA strand a little away from the centre of the palindromic site.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both statement I and Statement II are correct  
 (2) Both statement I and Statement II are incorrect  
 (3) Statement I is correct but Statement II is incorrect  
 (4) Statement I is incorrect but Statement II is correct

184. At which stage of life the oogenesis process is initiated?

- (1) Puberty    (2) Embryonic development stage  
 (3) Birth    (4) Adult

185. Identify the asexual reproductive structure associated with *Penicillium*:

- (1) Zoospores              (2) Conidia              (3) Gemmules              (4) Buds

## SECTION - B (BIOLOGY : ZOOLOGY)

186. Which of the following is not a desirable feature of a cloning vector?

- (1) Presence of origin of replication                      (2) Presence of a marker gene  
(3) Presence of single restriction enzymes site        (4) Presence of two or more recognition sites

187. The recombination frequency between the genes a & c is 5%, b & c is 15%, b & d is 9%, a & b is 20%, c & d is 24% and a & d is 29%. What will be the sequence of these genes on a linear chromosome?

- (1) a, d, b, c                      (2) d, b, a, c                      (3) a, b, c, d                      (4) a, c, b, d

188. Match List -I with List -II

	List-I (Biological Molecules)		List-II (Biological functions)
(a)	Glycogen	(i)	Hormone
(b)	Globulin	(ii)	Biocatalyst
(c)	Steroids	(iii)	Antibody
(d)	Thrombin	(iv)	Storage product

- (1) (a) – (iii), (b) – (ii), (c) – (iv), (d) – (i)                      (2) (a) – (iv), (b) – (ii), (c) – (i), (d) – (iii)  
(3) (a) – (ii), (b) – (iv), (c) – (iii), (d) – (i)                      (4) (a) – (iv), (b) – (iii), (c) – (i), (d) – (ii)

189. Select the incorrect statement regarding synapses.

- (1) The membranes of presynaptic and postsynaptic neurons are in close proximity in an electrical synapse  
(2) Electrical current can flow directly from one neuron into the other across the electrical synapse  
(3) Chemical synapses use neurotransmitters  
(4) Impulse transmission across a chemical synapse is always faster than that across an electrical synapse

190. Which one of the following statements is correct?

- (1) The atrio-ventricular node (AVN) generates an action potential to stimulate atrial contraction  
(2) The tricuspid and the bicuspid valves open due to the pressure exerted by the simultaneous contraction of the atria  
(3) Blood moves freely from atrium to the ventricle during joint diastole  
(4) Increased ventricular pressure causes closing of the semilunar valves.

191. Match List -I with List -II

	List-I		List-II
(a)	Bronchioles	(i)	Dense Regular connective Tissue
(b)	Goblet cell	(ii)	Loose Connective Tissue
(c)	Tendons	(iii)	Glandular Tissue
(d)	Adipose Tissue	(iv)	Ciliated Epithelium

- (1) (a) – (iv), (b) – (iii), (c) – (i), (d) – (ii)                      (2) (a) – (i), (b) – (ii), (c) – (iii), (d) – (iv)  
(3) (a) – (ii), (b) – (i), (c) – (iv), (d) – (iii)                      (4) (a) – (iii), (b) – (iv), (c) – (ii), (d) – (i)

192. Which of the following statements is not true?

- (1) Analogous structures are a result of convergent evolution  
(2) Sweet potato and potato is an example of analogy  
(3) Homology indicates common ancestry  
(4) Flippers of penguins and dolphins are a pair of homologous organs

193. Which of the following is a correct statement?
- (1) Cyanobacteria are a group of autotrophic organisms classified under Kingdom Monera.
  - (2) Bacteria are exclusively heterotrophic organisms
  - (3) Slime moulds are saprophytic organisms classified under Kingdom Monera.
  - (4) Mycoplasma have DNA, Ribosome and cell wall

194. Match List -I with List -II with respect to methods of Contraception and their respective actions.

	List-I		List-II
(a)	Diaphragms	(i)	Inhibit ovulation and Implantation
(b)	Contraceptive Pills	(ii)	Increase phagocytosis of sperm within Uterus
(c)	Intra Uterine Devices	(iii)	Absence of Menstrual cycle and ovulation following parturition
(d)	Lactational Amenorrhoea	(iv)	They cover the cervix blocking the entry of sperms

Choose the correct answer from the options given below:

- (1) (a) – (iv), (b) – (i), (c) – (iii), (d) – (ii)
  - (2) (a) – (iv), (b) – (i), (c) – (ii), (d) – (iii)
  - (3) (a) – (ii), (b) – (iv), (c) – (i), (d) – (iii)
  - (4) (a) – (iii), (b) – (ii), (c) – (i), (d) – (iv)
195. Ten E.coli cells with  $^{15}\text{N}$ -dsDNA are incubated in medium containing  $^{14}\text{N}$  nucleotide. After 60 minutes how many E.Coli cells will have DNA totally free from  $^{15}\text{N}$ ?
- (1) 20 cells
  - (2) 40 cells
  - (3) 60 cells
  - (4) 80 cells
196. Select the incorrect statement with respect to acquired immunity.
- (1) Primary response is produced when our body encounters a pathogen for the first time.
  - (2) Anamnestic response is elicited on subsequent encounters with the same pathogen.
  - (3) Anamnestic response is due to memory of first encounter.
  - (4) Acquired immunity is non-specific type of defense present at the time of birth
197. Statements related to human Insulin are given below. Which statement(s) is/are correct about genetically engineered Insulin?
- (a) Pro-hormone insulin contain extra stretch of C-peptide
  - (b) A-peptide and B-peptide Chains of insulin were produced separately in E.coli, extracted and combined by creating disulphide bond between them.
  - (c) Insulin used for treating Diabetes was extracted from Cattles and Pigs.
  - (d) Pro-hormone Insulin needs to be processed for converting into a mature and functional hormone.
- Choose the most appropriate answer from the options given below:
- (1) (a), (b) and (d) only
  - (2) (b) only
  - (3) (c) and (d) only
  - (4) (c), (d) and (e) only
198. If a colour blind female marries a man whose mother was also colour blind, what are the chances of her progeny having colour blindness?
- (1) 25%
  - (2) 50%
  - (3) 75%
  - (4) 100%

199. Which of the following are not the effects of Parathyroid hormone?

- (a) Stimulates the process of bone resorption
- (b) Decreases  $\text{Ca}^{2+}$  level in blood
- (c) Reabsorption of  $\text{Ca}^{2+}$  by renal tubules
- (d) Decreases the absorption of  $\text{Ca}^{2+}$  from digested food
- (e) Increases metabolism of carbohydrates

Choose the most appropriate answer from the options given below:

- (1) (a) and (c) only
- (2) (b), (d) and (e) only
- (3) (a) and (e) only
- (4) (b) and (c) only

200. Given below are two statements:

Statement I: In a scrubber the exhaust from the thermal plant is passed through the electric wires to charge the dust particles

Statement II: Particulate matter (PM 2.5) can not be removed by scrubber but can be removed by an electrostatic precipitator.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both statement I and Statement II are correct
- (2) Both statement I and Statement II are incorrect
- (3) Statement I is correct but Statement II is incorrect
- (4) Statement I is incorrect but Statement II is correct

NEET : 2022 - Paper Code

**Q5**

**ANSWER KEY**

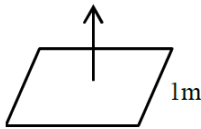
PHYSICS		CHEMISTRY		BIOLOGY			
Q 1	2	Q 51	2	Q 101	1	Q 151	3
Q 2	4	Q 52	1	Q 102	3	Q 152	3
Q 3	4*	Q 53	1	Q 103	2	Q 153	4
Q 4	2	Q 54	4	Q 104	3	Q 154	2
Q 5	3	Q 55	3	Q 105	3	Q 155	3
Q 6	3	Q 56	2	Q 106	1	Q 156	2
Q 7	2	Q 57	4	Q 107	2	Q 157	3
Q 8	4	Q 58	3	Q 108	3	Q 158	2
Q 9	3	Q 59	2	Q 109	4	Q 159	1
Q 10	4	Q 60	4	Q 110	3	Q 160	3
Q 11	3	Q 61	3	Q 111	1	Q 161	1
Q 12	3	Q 62	3	Q 112	2	Q 162	3
Q 13	1	Q 63	2	Q 113	4	Q 163	1
Q 14	1	Q 64	2	Q 114	2	Q 164	3
Q 15	3	Q 65	4	Q 115	4	Q 165	3
Q 16	2	Q 66	1	Q 116	3	Q 166	4
Q 17	4	Q 67	3	Q 117	3	Q 167	2
Q 18	3	Q 68	1	Q 118	3	Q 168	4
Q 19	4	Q 69	4	Q 119	2	Q 169	4
Q 20	2	Q 70	4	Q 120	1	Q 170	2
Q 21	3	Q 71	3	Q 121	2	Q 171	1
Q 22	2	Q 72	3	Q 122	2	Q 172	4
Q 23	4	Q 73	2	Q 123	2	Q 173	1
Q 24	3	Q 74	4	Q 124	3	Q 174	1
Q 25	4	Q 75	1	Q 125	3	Q 175	1
Q 26	3	Q 76	2	Q 126	1	Q 176	1
Q 27	2	Q 77	3	Q 127	2	Q 177	4
Q 28	2	Q 78	3	Q 128	3	Q 178	2
Q 29	4	Q 79	4	Q 129	3*	Q 179	4
Q 30	3	Q 80	1	Q 130	2	Q 180	3
Q 31	2	Q 81	4	Q 131	3	Q 181	4
Q 32	2	Q 82	4	Q 132	3	Q 182	2
Q 33	3	Q 83	4	Q 133	3	Q 183	1
Q 34	1	Q 84	2	Q 134	2	Q 184	2
Q 35	3	Q 85	4	Q 135	1	Q 185	2
Q 36	4	Q 86	2	Q 136	2	Q 186	3
Q 37	3	Q 87	2	Q 137	2	Q 187	4
Q 38	2	Q 88	1	Q 138	2	Q 188	4
Q 39	3	Q 89	4	Q 139	1	Q 189	4
Q 40	3	Q 90	2	Q 140	3	Q 190	3
Q 41	2	Q 91	3	Q 141	3	Q 191	1
Q 42	2	Q 92	3	Q 142	1	Q 192	4
Q 43	2	Q 93	1	Q 143	2	Q 193	1
Q 44	3	Q 94	1	Q 144	3	Q 194	2
Q 45	2	Q 95	1	Q 145	2	Q 195	3
Q 46	1	Q 96	4	Q 146	3	Q 196	4
Q 47	4	Q 97	1	Q 147	2	Q 197	2
Q 48	2	Q 98	3	Q 148	2	Q 198	4
Q 49	3	Q 99	1	Q 149	2	Q 199	2
Q 50	3	Q 100	3	Q 150	1	Q 200	2

## DETAILED SOLUTION

## Section-A (PHYSICS)

1 (2)

$$B = 0.5 \text{ T}$$



Angle between  $\vec{B}$  &  $\vec{A}$  is zero

$$\begin{aligned}\phi &= B \cdot A \cdot \cos 0 \\ &= 0.5 \times (1) \times 1\end{aligned}$$

$$= 0.5 \text{ Wb}$$

2 (4)

$$n = \sqrt{\epsilon_r \mu_r}$$

$$n = \frac{c}{v} \Rightarrow v = \frac{c}{n}$$

$$v = \left( \frac{c}{\sqrt{\epsilon_r \mu_r}} \right)$$

3 (bonus)

Using the equation

$$eV = h\nu - \phi$$

$$\text{or } eV = h\nu - h\nu_{Th}$$

$$\frac{eV_s}{2} = \frac{h\nu}{2} - h\nu_{Th} \quad \dots(1)$$

$$eV_s = h\nu - h\nu_{Th} \quad \dots(2)$$

Data Incorrect

4 (2)

Initially speed is zero, then increases & after some time it becomes constant.

5 (3)

$$d\vec{B} = \frac{\mu_0 (I d\vec{\ell} \times \vec{r})}{4\pi r^3}$$

The expression for magnetic field depends on current carrying element  $I d\vec{\ell}$ , which is a vector quantity, therefore, statement-I is correct and statement-II is wrong.

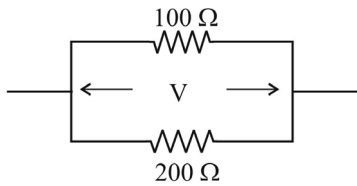
6 (3)

For conductors  $\alpha$  is (+ve)

For semiconductors  $\alpha$  is (-ve)



7 (2)



$$P = \frac{V^2}{R} \Rightarrow P \propto \frac{1}{R}$$

$$\frac{P_1}{P_2} = \frac{R_2}{R_1} = \frac{200}{100} = \frac{2}{1}$$

$$= 2 : 1$$

8 (4)

In (a) & (c) circuits, both the junctions are in same biasing conditions so offers equal resistances.

9 (3)

Peak voltage is  $\sqrt{2}$  times rms voltages in ac.

10 (4)

Slope of (x - t) graph, gives velocity

$$\frac{V_1}{V_2} = \frac{\tan \theta_1}{\tan \theta_2} = \frac{\tan 30^\circ}{\tan 45^\circ} = \frac{1}{\sqrt{3}}$$

11 (3)

Electric field is always perpendicular to equipotential surface.

12 (3)

$[MLT^{-2}A^{-2}]$  = Magnetic permeability

13 (1)

$$P = P_0 + \frac{4T}{R}$$

on expansion R increase, as a result, P decreases.

14 (1)

$$E = P \times t = 100 \times 10^3 \times 3600$$

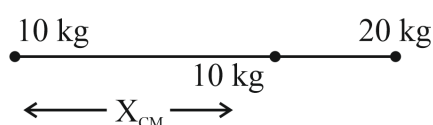
$$= 36 \times 10^7 \text{ J}$$

15 (3)

In half wave rectification

$$\Rightarrow f_{\text{output}} = 60 \text{ Hz}$$

16 (2)



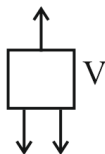
$$X_{CM} = \frac{20 \times 10}{20 + 10} = \frac{20}{3} \text{ m}$$

17 (4)

- (a) Radio wave (ii)  $\approx 102 \text{ m}$  (ii)  
 (b) Microwave  $\approx$  (iii)  $10.2 \text{ m}$  (iii)  
 (c) Infrared radiations  $\approx$  (iv)  $10.4 \text{ m}$  (iv)  
 (d) X-ray (i)  $\approx \text{\AA} = 10.10 \text{ m}$  (i)  
 (a) – (ii), (b) – (iii), (c) – (iv), (d) – (i)

18 (3)

$$\begin{aligned} \Rightarrow T &= W + f \\ &= 20000 + 3000 \\ &= 23000 \text{ N} \end{aligned}$$



$$\begin{aligned} \Rightarrow \text{Power} &= Tv \\ &= 23000 \times 1.5 \\ &= 34500 \text{ watts} \end{aligned}$$

19 (4)

$$y = (n\lambda) \left( \frac{D}{d} \right)$$

$$n_1 \lambda_1 = n_2 \lambda_2$$

$$(8)(600 \text{ nm}) = n_2 (400 \text{ nm})$$

$$n_2 = 12$$

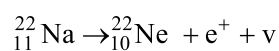
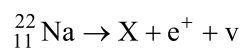
20 (2)

$$V = \frac{1}{4\pi \epsilon_0} \cdot \frac{Q}{R}$$

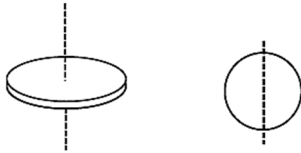
$$\therefore V \propto \frac{1}{R}$$

$\therefore$  Potential is more on smaller sphere.

21 (3)



22 (2)



$$k = \sqrt{\frac{1}{m}}$$

$$\Rightarrow \frac{k_1}{k_2} = \sqrt{\frac{I_1}{I_2}} = \sqrt{\frac{mR^2/2}{mR^2/4}} = \sqrt{2} : 1$$

23 (4)

First excited state  $\Rightarrow n = 2$

$$T_1 = 13.6 \frac{z^2}{n^2} = -\frac{13.6}{4} \text{ eV}$$

Second excited state  $\Rightarrow n = 3$

$$T_2 = 13.6 \frac{z^2}{n^2} = -\frac{13.6}{9} \text{ eV}$$

$$T_1 : T_2 = \frac{1}{4} : \frac{1}{9} = 9 : 4$$

24 (3)

Because angle of incidence is Brewster's angle so that angle between reflected and refracted ray is  $90^\circ$

$$\tan i_p = \mu = \sqrt{3}$$

$$i_p = 60^\circ = i$$

25 (4)

$$J = \sigma E \Rightarrow \frac{E}{\rho} = \frac{E\ell}{RA} = \frac{10 \times 10 \times \pi}{10 \times 10^{-4}}$$

$$\Rightarrow 10^5 \text{ A/m}^2$$

26 (3)

$$P = \frac{1}{f} = (\mu - 1) \left( \frac{1}{R_1} - \frac{1}{R_2} \right)$$

$$P = \left( \frac{3}{2} - 1 \right) \left( \frac{1}{0.2} + \frac{1}{0.2} \right)$$

$$P = \frac{1}{2} \left( \frac{2}{0.2} \right) = \frac{10}{2} = +5D$$

27 (2)

$$B = \mu_0 ni = \mu_0 \frac{N}{\ell} i$$

$$\therefore B = 4\pi \times 10^{-7} \times \frac{100}{10^{-3}} \times 1 = 12.56 \times 10^{-2} \text{ T}$$

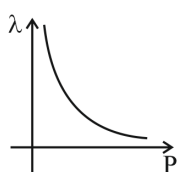
28 (2)

$$I_g = \frac{F}{m} = \frac{3}{60 \times 10^{-3}} = 50 \text{ N/kg}$$

29 (4)

$$\lambda = \frac{h}{p}$$

Graph will be hyperbolic



30 (3)

$$S_{nth} = u + \frac{a}{2}(2n-1)$$

$$S_{nth} \propto (2n-1)$$

$$= 1:3:5:7$$

31 (2)

$$\omega = \omega_0 + \alpha t$$

$$\alpha = \frac{\omega - \omega_0}{t}$$

$$= \frac{(3120 - 1200)}{16} \text{ rpm}$$

$$= \frac{1920}{16} \times \frac{2\pi}{60} \text{ rad/s}^2$$

$$= 4\pi \text{ rad/s}^2$$

32 (2)

Adiabatic curve is more steeper than isothermal.

33 (3)

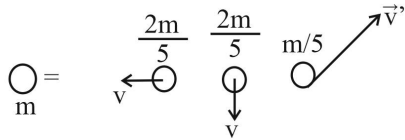
$$v \propto \sqrt{T}$$

$$\frac{v_i}{v_f} = \sqrt{\frac{1}{2}} = \frac{1}{\sqrt{2}}$$

34 (1)

Plane angle and solid angle are dimensionless but have units.

35 (3)



By conservation of momentum:

$$m(0) = \frac{2m}{5}(-v\hat{i}) + \frac{2m}{5}(-v\hat{j}) + \frac{m}{5}\vec{v}'$$

$$\Rightarrow \vec{v}' = 2v\hat{i} + 2v\hat{j}$$

$$\Rightarrow v' = \sqrt{(2v)^2 + (2v)^2}$$

$$= 2\sqrt{2}v$$

**Section-B (PHYSICS)**

36 (4)

$$\begin{aligned} \text{Area} &= \text{Length} \times \text{Breadth} \\ &= 55.3 \times 25 \\ &= 1382.5 \\ &= 14 \times 10^2 \end{aligned}$$

37 (3)

$$i_{\max} = \frac{E_{\max}}{R} = \frac{NBA\omega}{R}$$

$$i_{\max} = \frac{1000 \times 2 \times 10^{-5} \times \pi(10^2) \times 2}{12.56}$$

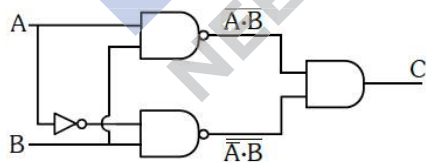
$$i_{\max} = 1A$$

38 (2)

$$E = \frac{Kp}{r^3} \sqrt{1 + 3\cos^2 \theta}$$

$$\therefore E \propto \frac{1}{r^3}$$

39 (3)



$$C = \overline{A \cdot B} + \overline{\overline{A} \cdot \overline{B}}$$

Using De-Morgan Theorem

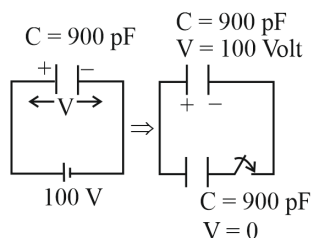
$$C = \overline{A \cdot B} + \overline{\overline{A} \cdot \overline{B}}$$

$$C = \overline{B(A + \overline{A})} = \overline{B}$$

Therefore

A	B	C
0	0	1
0	1	0
1	0	1
1	1	0

40 (3)



Common potential

$$V_c = \frac{C_1 V_1 + C_2 V_2}{C_1 + C_2}$$

$$= \frac{C \times 100 + C \times 0}{C + C}$$

$$= 50 \text{ Volt}$$

Electrostatic energy stored

$$= 2 \times \frac{1}{2} CV^2 = CV^2$$

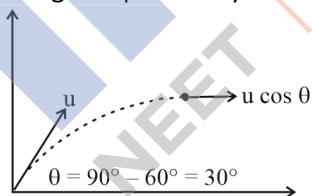
$$= 2.25 \times 10^{-6} \text{ J}$$

41 (2)

$$i_c = \sin^{-1} \left( \frac{3}{4} \right)$$

42 (2)

At highest point only horizontal component of velocity remains  $\Rightarrow u_x = u \cos \theta$



$$u_x = u \cos \theta = 10 \cos 30^\circ$$

$$= 5\sqrt{3} \text{ ms}^{-1}$$

43 (2)

Resistance of P & Q should be approx. equal as it decreases error in experiment.

44 (3)

$$Y_{\text{copper}} < Y_{\text{steel}}$$

45 (2)

$$\omega = 100$$

$$v = \frac{\omega}{2\pi} = \frac{100}{2\pi} = \frac{50}{\pi} \text{ Hz}$$

Resonance frequency

$$v_0 = \frac{1}{2\pi\sqrt{LC}} = \frac{1}{2\pi\sqrt{10 \times 10 \times 10^{-6}}}$$

$$= \frac{50}{\pi} \text{ Hz}$$

46 (1)

$$(n) (1.1) = (n + 1)$$

$$0.1(n) = 1$$

$$n = 10$$

$$\text{Required number of oscillation} = n + 1 = 10 + 1 = 11$$

47 (4)

$$V = (\text{no. of moles})(22.4 \text{ litre})$$

$$= \frac{\text{mass}}{\text{molar mass}} (22.4 \times 10^{-3} \text{ m}^3)$$

$$= \frac{4.5 \times 10^3}{18} \times 22.4 \times 10^{-3} \text{ m}^3$$

$$= 5.6 \text{ m}^3$$

48 (2)

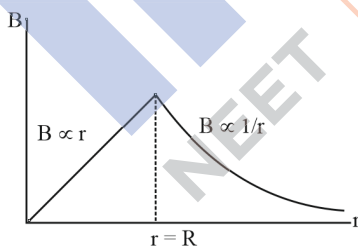
$$\text{Gravitational constant} = [M^{-1}L^3T^{-2}]$$

$$\text{Gravitational potential energy} = [ML^2T^{-2}]$$

$$\text{Gravitational potential} = [L^2T^{-2}]$$

$$\text{Gravitational intensity} = [LT^{-2}]$$

49 (3)



50 (3)

$$R = R_0 (A)^{1/3}$$

$$\frac{R_1}{R_2} = \left( \frac{125}{64} \right)^{1/3} = \frac{5}{4}$$

## SECTION - A (CHEMISTRY)

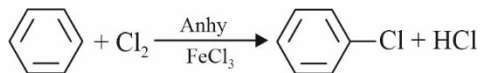
51. (2)

The oxidation state of K in  $\text{KO}_2$  is (+1)

52. (1)

IUPAC name of element with atomic number 119 is ununennium.

53. (1)



54. (4)

Li is used in making batteries, liquid Na is a coolant in nuclear reactors. Cs is used in photoelectric cells and KOH as a base is used to absorb acidic oxides like  $\text{CO}_2$ .

55. (3)

I – Cl bond is stronger than I-I bond due to additional electrostatic force between I and Cl.

56. (2)

Both statements are correct independently but the reason is not the correct reason of the assertion.

57. (4)

As compared to hydrocarbons of similar mass aldehydes and ketones will have greater dipole – dipole interactions.

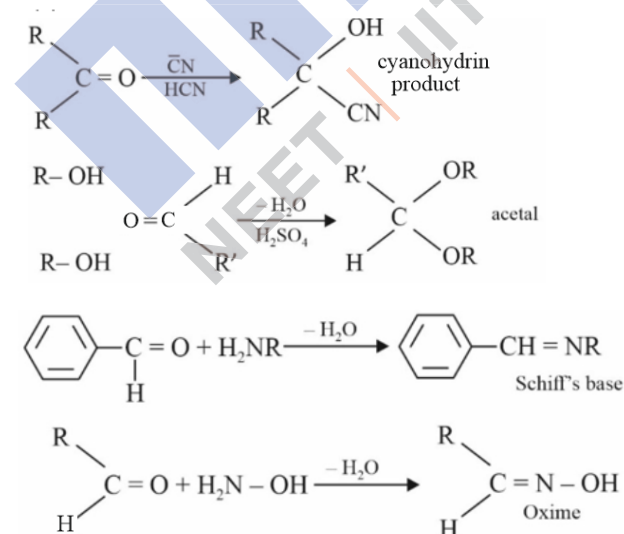
58. (3)

Carbon atoms in diamond are  $\text{sp}^3$  hybridized and those in graphite are  $\text{sp}^2$  hybridized

59. (2)

Cimetidine is an antacid. Seldane is an antihistaminic drug.

60. (4)



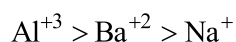
61. (3)

Aliphatic diazonium salts are not stable Aromatic diazonium salts exist at low temp of  $0 - 4^\circ\text{C}$ .

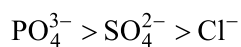
62. (3)



For the coagulation of a negative sol flocculating power is



But for coagulating a positive sol the flocculating power will be



63. (2)

The boiling point order is:  $\text{H}_2\text{S} < \text{H}_2\text{Se} < \text{H}_2\text{Te} < \text{H}_2\text{O}$

64. (2)

$$\text{Molality} = \frac{\text{moles of solute}}{\text{mass of solute(kgs)}} \Rightarrow 1 = \frac{0.5}{W(\text{kgs})}$$

$$W_{\text{solvent}}(\text{kgs}) = 0.5 = 500\text{g}$$

65. (4)

In diborane the boron atoms are  $\text{sp}^3$  hybridized

66. (1)

$\text{MgH}_2$  is an ionic or saline hydride,  $\text{GeH}_4$  is an electron precise hydride with 8 electrons around Ge,

$\text{B}_2\text{H}_6$  is an electron deficient hydride and HF is electron rich.

67. (3)

Enantiomers are non-superimposable mirror images of each other.

68. (1)

$$\text{p}^{\text{H}} = \text{p}^{\text{ka}} + \log \frac{[\text{CH}_3\text{COONa}]}{[\text{CH}_3\text{COOH}]}$$

$$\text{p}^{\text{H}} = 4.57 + \log \frac{0.1}{0.01} = 5.57$$

69. (4)

$[\text{Ag}(\text{H}_2\text{O})_2]^+[\text{Ag}(\text{CN})_2]^-$  is called diaquasilver(I)dicyanidoargentate(I)

70. (4)

Cyclohepta-1, 3, 5-triene is not aromatic as one carbon is saturated ( $\text{sp}^3$ )

71. (3)

Kjeldahl's method does not work for heterocyclic and azo compounds.

72. (3)

Enzymes are proteins

73. (2)

Gadolinium has outer configuration of  $[\text{Xe}]4f^7 5d^1 6s^2$

Its third ionization energy is low due to highly exchange energy and hence stability of the half-filled f-subshell.

74. (4)

$O_2^+(15)$  will have configuration  $\sigma_{1s}^2 \sigma_{1s}^{*2} \sigma_{2s}^2 \sigma_{2s}^{*2} \sigma_{2p_x}^2 \left\{ \begin{array}{l} \pi_{2p_y}^2 \\ \pi_{2p_z}^2 \end{array} \right\} \left\{ \begin{array}{l} \pi_{2p_y}^{*1} \\ \pi_{2p_z}^{*1} \end{array} \right\}$ . This ion is paramagnetic

75. (1)

$E_{\text{cell}}^0 = E_{\text{cathode}}^0 - E_{\text{anode}}^0 = (+1.510) - (1.229) = +0.287V$ . As  $E_{\text{cell}}^0$  is positive the cell will work

76. (2)

Maximum work is done in the case where area under the curve is maximum

77. (3)

At different positions  $-NO_2$  affects acidic strength differently. The order of acidic strength is p-nitro phenol > o-nitrophenol > m-nitrophenol

78. (3)

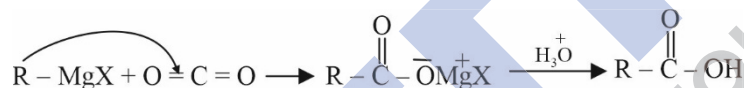
For a zero-order reaction Rate Vs conc graph will be straight line parallel to the x-axis

For a 1<sup>st</sup> order reaction  $t_{\frac{1}{2}}$  vs concentration will again be a straight line parallel to x-axis.

79. (4)

Shapes of  $d_{x^2-y^2}$  and  $d_{z^2}$  are not similar to each other.

80. (1)



81. (4)

Maximum lone pair – lone pair repulsions are there in  $XeF_2$

82. (4)

$p_i = p_i \chi_i$  is not a correct form of Dalton's law of partial pressures

83. (4)

Thermosetting polymers such as Bakelite cannot be moulded again and are hence not reusable.

84. (2)

$$\text{Moles of HCl} = \frac{50}{1000} \times 0.5 = 0.025 \text{ moles}$$

$$\text{So, moles of CaCO}_3 \text{ used} = \frac{0.025}{2} = 0.0125 \text{ moles} = 1.25 \text{ g}$$

$$95\% (\text{Total mass of CaCO}_3) = 1.25 \text{ g}$$

$$\text{Total mass of CaCO}_3 = \frac{1.25}{0.95} = 1.32 \text{ g}$$

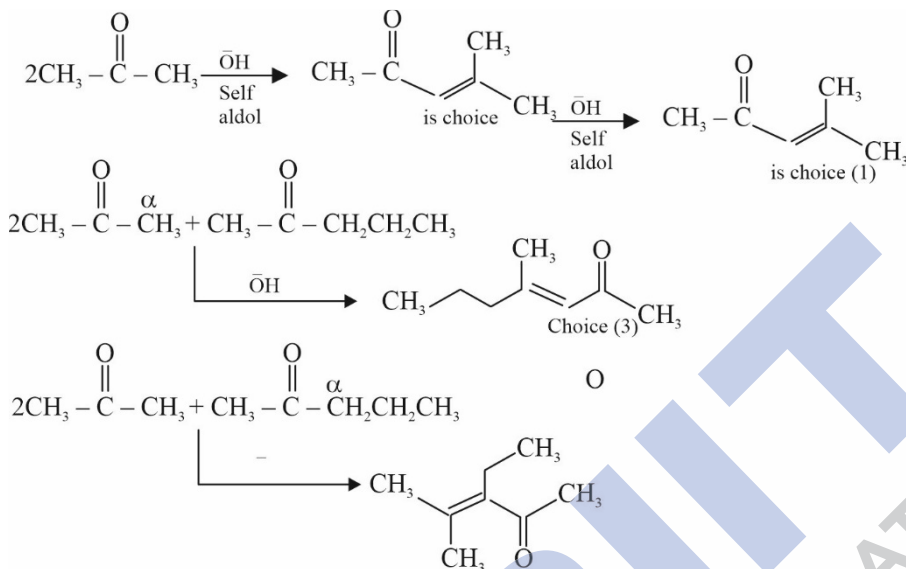
85. (4)

Silver will not be able to reduce  $Cu^{+2}$  to Cu.

SECTION - B (CHEMISTRY)

86. (2)

When acetone & 2-pentanone are added in a base for an aldol condensation reaction there are several possibilities that may arise for product formation.



But structure in choice (2) is not possible. It will be possible with 3-pentanone.

87. (2)

$$k = \frac{2.303}{5} \log \frac{0.1}{0.001} = \frac{2.303 \times 2}{5} = \frac{4.606}{5} = 0.921 \text{ min}^{-1}$$

88. (1)



$$3 \times 10^{-59} = k_c = \frac{[\text{O}_3]^2}{[\text{O}_2]^3} = \frac{[\text{O}_3]^2}{[4 \times 10^{-2}]^3}$$

$$[\text{O}_3]^2 = 192 \times 10^{-65} \Rightarrow [\text{O}_3] = 4.38 \times 10^{-32}$$

89. (4)



90. (2)

Hematite:  $\text{Fe}_2\text{O}_3$

Magnetite:  $\text{Fe}_3\text{O}_4$

Calamine:  $\text{ZnCO}_3$

Kaolinite:  $[\text{Al}_2(\text{OH})_4\text{Si}_2\text{O}_5]$

91. (3)

Statement I is true but statement II is incorrect as 3°-alcohols are the most reactive and give immediate turbidity with Luca's reagent.

92. (3)

$$E_{\text{cell}} = 1.05 - \frac{0.059}{2} \log \frac{[\text{Ni}^{+2}]}{[\text{Ag}^+]^2}$$

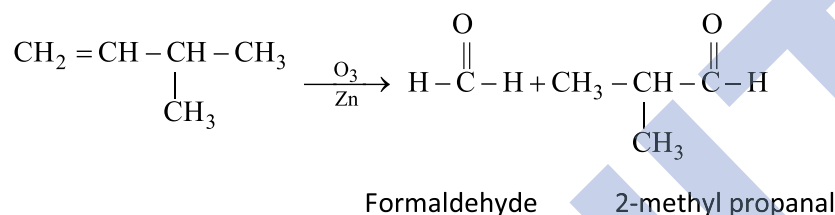
$$= 1.05 - \frac{0.059}{2} \log \frac{[10^{-3}]}{[10^{-3}]^2}$$

$$= 1.05 - \frac{0.059 \times 3}{2} = 1.05 - 0.0885 = 0.9615 \text{ volt}$$

**PLEASE NOTE**

There is a misprint in the question. The  $E_{\text{cell}}^0$  is incorrectly given as 10.5 V. This should have been 1.05 volt.

93. (1)



94. (1)

Oxidation state in neutral medium changes from +7 to +4 for  $\text{MnO}_4^-$  to  $\text{MnO}_2$

95. (1)

$$\frac{a^3 \times d}{M} \times N_A = Z \Rightarrow \frac{(3.608 \times 10^{-8})^{3 \times 8.92}}{M} \times 6 \times 10^{23} = 4$$

$$M = \frac{46.96 \times 10^{-24} \times 8.92 \times 6.022 \times 10^{23}}{4} = 63 \text{ g/mole}$$

Closest answer is choice (1)

96. (4)

$$P = \frac{nRT}{V} = \frac{2 \times 0.0381 \times 300}{10} = 4.986 \text{ bar}$$

97. (1)

$$105.8 \times \frac{2}{4} = 52.9 \text{ pm} \Rightarrow r_{\text{Li}^{+2}} = 52.9 \times \frac{3^2}{3} = 158.7 \text{ pm}$$

98. (3)

More the number of strong field ligands in a complex greater the energy absorbed by a complex

99. (1)

The IUPAC name giving the lowest number to the functional group obeying the lowest set of local is

1-bromo-5-chloro-4-methylhexan-3-ol

100. (3)

Pollution due to oxides of sulphur gets enhanced due to hydrocarbons and oxidizing agents like  $\text{O}_3$  &  $\text{H}_2\text{O}_2$